

The syllabus is in tune with the syllabus adapted by the Telangana State Board of Intermediate Education (TSBIE), for Intermediate course with effect from the academic year 2019-2020 (1st year) (100%) and 2020-2021 (2nd year) (70%) and is designed at the level of Intermediate Course and equivalent to 10 + 2 (10 plus 2) scheme of Examination conducted by Telangana State Board of Intermediate Education. The syllabus is designed to indicate the scope of subjects included for TS EAMCET-2021. The topics mentioned therein are not to be regarded as exhaustive. Questions may be asked in TS EAMCET-2021 to test the student's knowledge and intelligent understanding of the subject. The syllabus is applicable to students of both the current and previous batches of Intermediate Course, who desire to appear for TS EAMCET-2021.

BOTANY

1) DIVERSITY IN THE LIVING WORLD:

The living world: What is living? Diversity in the living world; Taxonomic categories and taxonomical aids.

Biological Classification: Five kingdom classification - Monera, Protista, Fungi, Plantae and Animalia, Three domains of life (six kingdom classification), Viruses, Viroids, Prions & Lichens.

Science of plants – Botany: Origin, Development, Scope of Botany and Branches of Botany.

Plant Kingdom: Salient features, classification and alternation of generations of the plants of the following groups - Algae, Bryophytes, Pteridophytes, Gymnosperms and Angiosperms.

2) STRUCTURAL ORGANISATION IN PLANTS- MORPHOLOGY:

Morphology of flowering Plants

Vegetative: Parts of a typical Angiospermic plant; Vegetative morphology and modifications- Root, Stem and Leaf- types; Venation, Phyllotaxy.

Reproductive: Inflorescence - Racemose, Cymose and special types (in brief).

Flower: Parts of a flower and their detailed description; Aestivation, Placentation.

Fruits: Types- True, False and Parthenocarpic fruits.

3) REPRODUCTION IN PLANTS:

Modes of Reproduction: Asexual reproduction, binary fission, Sporulation, budding, fragmentation, vegetative propagation in plants, Sexual reproduction, Overview of angiosperm life cycle.

Sexual Reproduction in Flowering Plants: Stamen, microsporangium, pollen grain. Pistil, megasporangium (ovule) and embryo sac; Development of male and female gametophytes. Pollination - Types, agents, Out breeding devices and Pollen - Pistil interaction. Double Fertilization; Post fertilisation events: Development of endosperm and embryo; development of seed, Structure of Dicotyledonous and Monocotyledonous seeds, Significance of fruit and seed. Special modes - Apomixis, parthenocarpy, polyembryony.

4) PLANT SYSTEMATICS:

Taxonomy of angiosperms: Introduction. Types of Systems of classification (in brief). Semi-Technical description of a typical flowering plant. Description of Families: Fabaceae, Solanaceae and Liliaceae.

5) CELL STRUCTURE AND FUNCTION:

Cell - The Unit of Life: Cell- Cell theory and cell as the basic unit of life- overview of the cell. Prokaryotic and Eukaryotic cells, Ultra Structure of Plant cell (structure in detail and functions in brief), Cell membrane, Cell wall, Cell organelles: Endoplasmic reticulum, Mitochondria, Plastids, Ribosomes, Golgi bodies, Vacuoles, Lysosomes, Microbodies, Centrosome and Centriole, Cilia, Flagella, Cytoskeleton and Nucleus. Chromosomes: Number, structural organization; Nucleosome.

Biomolecules: Structure and function of Proteins, Carbohydrates, Lipids and Nucleic acids.

Cell cycle and Cell Division: Cell cycle, Mitosis, Meiosis - significance.

6) INTERNAL ORGANISATION OF PLANTS:

Histology and Anatomy of Flowering Plants: Tissues - Types, structure and functions: Meristematic; Permanent tissues - Simple and Complex tissues. Tissue systems - Types, structure and function: Epidermal, Ground and Vascular tissue systems. Anatomy of Dicotyledonous and Monocotyledonous plants - Root, Stem and Leaf. Secondary growth in Dicot stem and Dicot root.

7) PLANT ECOLOGY:

Ecological Adaptations, Succession and Ecological Services: Introduction. Plant communities and Ecological adaptations: Hydrophytes, Mesophytes and Xerophytes. Plant succession. Ecological services - Carbon fixation, Oxygen release and pollination.

8) PLANT PHYSIOLOGY:

Enzymes: Chemical Reactions, Enzymatic Conversions, Nature of Enzyme Action, Factors Affecting Enzyme Activity, Temperature and pH, Concentration of Substrate, Classification and Nomenclature of Enzymes, Co-factors.

Photosynthesis in Higher Plants: Early Experiments, Site of Photosynthesis, Pigments involved in Photosynthesis, Light Reaction, The Electron Transport-Splitting of Water, Cyclic and Noncyclic Photo-phosphorylation, Chemiosmotic Hypothesis, Biosynthetic phase-The Primary Acceptor of CO₂, The Calvin Cycle, The C₄ Pathway, Photorespiration, Factors affecting Photosynthesis.

Respiration of Plants: Glycolysis, Fermentation, Aerobic Respiration - Tricarboxylic Acid Cycle, Electron Transport System (ETS) and Oxidative Phosphorylation, The Respiratory Balance Sheet, Amphibolic Pathway, Respiratory Quotient.

Plant Growth and Development:

Plant Growth Regulators- Characteristics, Discovery, Physiological effects of Plant Growth Regulators, Auxins, Gibberellins, Cytokinins, Ethylene, Absciscic acid.

9) GENETICS:

Principles of Inheritance and Variation: Mendel's Experiments, Inheritance of one gene (Monohybrid Cross)-Back cross and Test cross, Law of Dominance, Law of Segregation or Law of purity of gametes, Deviations from Mendelian concept of dominance

- Incomplete Dominance, Co-dominance, Explanation of the concept of dominance, Inheritance of two genes- (Dihybrid Cross) Law of Independent Assortment, Chromosomal Theory of Inheritance, Linkage and Recombination, Mutations, Significance of mutations.

10) MOLECULAR BIOLOGY:

Molecular Basis of inheritance: The DNA- Structure of Polynucleotide Chain, Packaging of DNA Helix, The Search for Genetic Material, Transforming Principle, Biochemical Characterisation of Transforming Principle, The Genetic Material is DNA, Properties of Genetic Material (DNA versus RNA), RNA World, Replication - The Experimental Proof, The Machinery and the Enzymes, Transcription-Transcription Unit, Transcription Unit and the Gene, Types of RNA and the process of Transcription, Genetic Code-Mutations and Genetic Code, tRNA- the Adapter Molecule, Translation, Regulation of Gene Expression-The Lac operon.

11) Biotechnology:

Principles and processes of Biotechnology: Principles of Biotechnology-Construction of the first artificial recombinant DNA molecule, Tools of Recombinant DNA Technology- Restriction Enzymes, Cloning Vectors, Competent Host (For Transformation with Recombinant DNA), Processes of Recombinant DNA Technology- Isolation of the Genetic Material (DNA), Cutting of DNA at Specific Locations, Separation and isolation of DNA fragments, Insertion of isolated gene into a suitable vector, Amplification of Gene of Interest using PCR, Insertion of Recombinant DNA into the Host, Cell/Organism, Selection of Transformed host cells, Obtaining the Foreign Gene Product, Downstream Processing.

Biotechnology and its applications: Biotechnological Applications in Agriculture-Bt Cotton, Pest Resistant Plants, Other applications of Biotechnology - Insulin, Gene therapy, Molecular Diagnosis, ELISA, DNA fingerprinting, Transgenic plants, Bio-safety and Ethical issues- Biopiracy.

12) MICROBES AND HUMAN WELFARE:

Microbes in Human Welfare: Microbes in Household Products, Microbes in Industrial Products-Fermented Beverages, Antibiotics, Chemicals, Enzymes and other Bioactive Molecules, Microbes in Sewage Treatment, Primary treatment, Secondary treatment or Biological treatment, Microbes in Production of Biogas, Microbes as Biocontrol Agents, Biological control of pests and diseases, Microbes as Biofertilisers, Challenges posed by Microbes.

The following topics are the deleted content of Botany (30%) in the Intermediate second year for the academic year 2020-21 by the TSBIE as per Telugu academy text book. The same is deleted for the TS EAMCET-2021 examination.

Unit I - PLANT PHYSIOLOGY

Chapter 1: Transport in Plants

Chapter 2: Mineral Nutrition

Chapter 6: Plant Growth and Development

- 6.1 Growth
- 6.2 Differentiation, Dedifferentiation and Redifferentiation
- 6.3 Development
- 6.5 Seed Dormancy
- 6.6 Photoperiodism
- 6.7 Vernalisation

Unit II - MICROBIOLOGY

Chapter 7: Bacteria

Chapter 8: Viruses

Unit VI- PLANTS, MICROBES AND HUMAN WELFARE

Chapter 13: Strategies for enhancement in Food production

ZOOLOGY

1) ZOOLOGY - Diversity of Living World:

What is life?; Nature, Scope & meaning of zoology; Branches of Zoology; Need for classification- Zoos as tools for classification; Basic principles of Classification: Biological system of classification- (Phylogenetic classification only); Levels or Hierarchy of classification; Nomenclature - Bi & Trinominal; Species concept; Kingdom Animalia; Biodiversity - Meaning and distribution, Genetic diversity, Species diversity, Ecosystem diversity(alpha, beta and gamma), other attributes of biodiversity, role of biodiversity, threats to biodiversity, methods of conservation, IUCN Red data books, Conservation of wild life in India -Legislation, Preservation, Organisations, Threatened species.

2) STRUCTURAL ORGANIZATION IN ANIMALS:

Levels of organization, Multicellularity: Diploblastic & Triploblastic conditions; Asymmetry, Symmetry: Radial symmetry, and Bilateral symmetry. Acoelomates, Pseudocoelomates and Eucoelomates: Schizo & Entero coelomates (Brief account of formation of coelom); Tissues: Epithelial, Connective, Muscular and Nervous tissues.

3) ANIMAL DIVERSITY-I: INVERTEBRATE PHYLA:

General Characters – (Strictly restrict to 8 salient features only, Classification up to Classes with two or three examples – Brief account only). Porifera; Cnidaria; Ctenophora; Platyhelminthes; Nematoda; Annelida (Include Earthworm as a type study); Arthropoda; Mollusca; Echinodermata; Hemichordata.

4) ANIMAL DIVERSITY-II: PHYLUM: CHORDATA

General Characters –Classification up to Classes - (Brief account only with two or three examples). Phylum: Chordata; Sub phylum: Urochordata; Sub phylum: Cephalochordata; Sub phylum: Vertebrata; Super class: Agnatha, Class Cyclostomata; Super class: Gnathostomata, Super class pisces, Class: Chondrichthyes, Class: Osteichthyes; Tetrapoda, Class: Amphibia (Including the brief type study of Frog) Class: Reptilia, Class: Aves, Class: Mammalia.

5) LOCOMOTION & REPRODUCTION IN PROTOZOA:

Locomotion: Definition, types of locomotor structures pseudopodia (basic idea of pseudopodia without going into different types), flagella & cilia (Brief account giving two examples each); Flagellar & Ciliary movement - Effective & Recovery strokes in Euglena, Synchronal & Metachronal movements in Paramecium; Reproduction: Definition, types. Asexual Reproduction: Transverse binary fission in Paramecium & Longitudinal binary fission in Euglena. Multiple fission, Sexual Reproduction.

6) BIOLOGY & HUMAN WELFARE:

Parasitism and parasitic adaptation; Health and disease: introduction Life cycle, Pathogenicity, Treatment & Prevention (Brief account only) 1. Entamoeba histolytica 2. Plasmodium vivax 3. Ascaris lumbricoides 4. Wuchereria bancrofti; Brief account of pathogenicity, treatment & prevention of Typhoid, Pneumonia, Common cold, & Ring worm; Tobacco, Drugs and Alcohol abuse (TDA Abuse).

7) TYPE STUDY OF PERIPLANETA AMERICANA:

Habitat and habits; External features; Locomotion; Digestive system; Respiratory system; Circulatory system; Excretory system; Nervous system - sense organs, structure of ommatidium; Reproductive system.

8) ECOLOGY & ENVIRONMENT:

What is Ecology, Importance of Ecology, Organisms and Environment: Ecology, population, communities, habitat, niche, biome and ecosphere (definitions only); Ecosystem: Elementary aspects only, Abiotic factors- Light, Temperature & Water (Biological effects only), Ecological adaptations; Population interactions; Ecosystems: Types, Components, Lake ecosystem; Food chains, Food web, Productivity and Energy flow in Ecosystem, Ecological pyramids - Pyramids of numbers, biomass and energy; Nutrient cycling - Carbon, Nitrogen & Phosphorous cycles (Brief account); Population attributes: Growth, Natality and Mortality, Age distribution, Population regulation; Environmental issues.

9) HUMAN ANATOMY AND PHYSIOLOGY-I:

Breathing and Respiration: Respiratory organs in animals; Respiratory system in humans; Mechanism of breathing and its regulation in humans - Exchange of gases, transport of gases and regulation of respiratory movements, Respiratory volumes; Respiratory disorders: Asthma, Emphysema, Bronchitis, Pneumonia, Occupational respiratory disorders - Asbestosis, Silicosis, Siderosis, Black Lung Disease in coal mine workers.

10) HUMAN ANATOMY AND PHYSIOLOGY-II:

Body Fluids and Circulation: Lymphatic system; Clotting of blood; Human circulatory system - structure of human heart and blood vessels; Cardiac cycle, cardiac output, double circulation, circulatory pathways, Portal circulation and coronary circulation; regulation of cardiac activity; Disorders of circulatory system: Hypertension, coronary artery disease, angina pectoris, heart failure.

Excretory products and their elimination: Modes of excretion - Ammonotelism, Ureotelism, Uricotelism, Excretory organs; Human excretory system - structure of kidney and nephron; Urine formation, osmoregulation; Regulation of kidney function -Renin-Angiotensin - Aldosterone system, Atrial Natriuretic Factor, ADH and diabetes insipidus; Role of other organs in excretion; Disorders: Uraemia, renal failure, renal calculi, glomerular nephritis, dialysis using artificial kidney and kidney transplantation.

11) HUMAN ANATOMY AND PHYSIOLOGY-III:

Muscular and Skeletal system: The muscle - ultra structure; Contractile proteins & Mechanism of muscle contraction, muscle fatigue, cori cycle, Types of muscle fibers.

Neural control and co-ordination: Nervous system in human beings - Central nervous system, Peripheral nervous system and Somatic and autonomic neural system; Generation and conduction of nerve impulse, Synaptic transmission.

12) HUMAN ANATOMY AND PHYSIOLOGY-IV:

Endocrine system and chemical co-ordination, Endocrine glands and hormones; Human endocrine system - Hypothalamus, Pituitary, Pineal, Thyroid, Parathyroid, Thymus gland, Adrenal, Pancreas, Gonads; Mechanism of hormone action (Elementary idea only), hormones of kidney, heart and gastrointestinal tract, Role of hormones as messengers and regulators; Hypo and Hyper activity and related disorders: Common disorders -Dwarfism, acromegaly, cretinism, goiter, exophthalmic goiter, diabetes, Addison's disease, Cushing's syndrome (Diseases & disorders to be dealt in brief).

Immune system: Basic concepts of Immunology - Types of Immunity - Innate Immunity, Acquired Immunity, Active and Passive Immunity, Cell mediated Immunity and Humoral Immunity, Cells of immune system, organs of immune system, soluble mediators of immunity, vaccination or immunization and immunological disorders.

13) HUMAN REPRODUCTION:

Human Reproductive System: Male and female reproductive systems; Microscopic anatomy of testis & ovary; Gametogenesis, Spermatogenesis & Oogenesis; Menstrual cycle; Fertilization, Embryo development up to blastocyst formation, Implantation; Pregnancy, placenta formation, Parturition, Lactation (elementary idea).

Reproductive Health: Need for reproductive health and prevention of sexually transmitted diseases (STD); Birth control - Need and methods, contraception and medical termination of pregnancy (MTP); Amniocentesis; infertility and assisted reproductive technologies - IVF-ET, ZIFT, GIFT (elementary idea for general awareness).

14) GENETICS:

Heredity and variations : Mendel's laws of inheritance with reference to *Drosophila*. (*Drosophila melanogaster* Grey, Black body colour; long, vestigial wings), Pleiotropy; Multiple alleles:

Inheritance of blood groups and Rh-factor; Codominance (Blood groups as example); Elementary idea of polygenic inheritance; Skin colour in humans. Sex determination – in humans, birds, *Bombus* moth, genic balance theory of sex determination in *Drosophila melanogaster* and honey bees; Sex linked inheritance – Haemophilia, Colour blindness; Mendelian disorders in humans: Thalassaemia, Haemophilia, Sickle celled anaemia, cystic fibrosis PKU, Alkaptonuria; Chromosomal disorders –Down's syndrome, Turner's syndrome and Klinefelter syndrome; Genome, Human Genome Project and DNA Finger Printing.

15) APPLIED BIOLOGY:

Animal Husbandry, Dairy management, Animal breeding, Apiculture, Beekeeping, Poultry management, Pisciculture, Fishery management, Application of Biotechnology in medicine: Human insulin and vaccine production; Molecular diagnosis Gene Therapy; Transgenic animals; Vaccines, Cancer biology, stem cells.

Note: (Up to primary concepts in Applied Biology)

The following topics are the deleted content of Zoology (30%) in the Intermediate second year for the academic year 2020-21 by the TSBIE as per Telugu academy text book. The same is deleted for the TS EAMCET-2021 examination.

1) HUMAN ANATOMY AND PHYSIOLOGY-I

Unit-I A: Digestion and absorption:

1.1 Digestive System

- 1.1.1 Alimentary canal
- 1.1.2 Digestive glands
- 1.1.3 Physiology of digestion
- 1.1.4 Peristalsis
- 1.1.5 Absorption of digested food

- 1.1.6 Gastro intestinal hormones
- 1.1.7 Calorific values of carbohydrates, proteins and fats
- 1.1.8 Disorders of the digestive system

2) **HUMAN ANATOMY AND PHYSIOLOGY-III**

Unit-III A: Topics in Musculo-Skeletal System

3.2 The Skeleton

- 3.2.1 Axial skeleton
- 3.2.2 Appendicular skeleton

3.3 Joints

- 3.3.1 Types of joints
- 3.3.2 Structure of synovial joint.

3.4 Disorders of the muscular and skeletal system: myasthenia gravis, tetany, muscular dystrophy, arthritis, osteoporosis, gout.

3) **HUMAN ANATOMY AND PHYSIOLOGY-III**

Unit-III B : Topics in Neural control and coordination

3.7 Reflex action and reflex arc;

3.8 Sensory reception and processing

- 3.8.1 The Eye
- 3.8.2 The mechanism of vision
- 3.8.3 The Ear
- 3.8.4 The mechanism of hearing

4) **ORGANIC EVOLUTION**

Unit-VII:

7.1 Origin of Life

7.2 Biological Evolution

7.3 Evidences for Biological Evolution (palaeontological, comparative anatomical, embryological and molecular evidences);

7.4 Theories of evolution: Lamarckism (in brief), Darwin's theory of Evolution- Natural Selection with example (Kettlewell's experiments on Biston betularia), Mutation Theory of Hugo de Vries;

7.5 Modern synthetic theory of Evolution

7.6 Mechanism of evolution- Hardy Weinberg law, Evolutionary forces, Types of Natural Selection; Gene flow and genetic drift; Human evolution; Speciation - Allopatric, sympatric; Reproductive isolation.

7.7 A brief account of evolution

7.8 Origin and evolution of man

5) **APPLIED BIOLOGY**

Unit-VIII

8.12 Biomedical Technology: Diagnostic Imaging (X-ray, CT scan, MRI), ECG, EEG; ELISA;

PHYSICS

1) PHYSICAL WORLD: What is physics? Scope and excitement of Physics, Physics, technology and society, Fundamental forces in nature, Gravitational Force, Electromagnetic Force, Strong Nuclear Force, Weak Nuclear Force, Towards Unification of Forces, Nature of physical laws.

2) UNITS AND MEASUREMENTS: Introduction, The international system of units, Measurement of Length, Measurement of Large Distances, Estimation of Very Small Distances: Size of a Molecule, Range of Lengths, Measurement of Mass, Range of Masses, Measurement of time, Accuracy, precision of instruments and errors in measurement, Systematic errors, random errors, least count error, Absolute Error, Relative Error and Percentage Error, Combination of Errors, Significant figures, Rules for Arithmetic Operations with Significant Figures, Rounding off the Uncertain Digits, Rules for Determining the Uncertainty in the Results of Arithmetic Calculations, Dimensions of Physical Quantities, Dimensional Formulae and dimensional equations, Dimensional Analysis and its Applications, Checking the Dimensional Consistency of Equations, Deducing Relation among the Physical Quantities.

3) MOTION IN A STRAIGHT LINE: Introduction, Position, path length and displacement, Average velocity and average speed, Instantaneous velocity and speed, Acceleration, Kinematic equations for uniformly accelerated motion, Relative velocity.

4) MOTION IN A PLANE: Introduction, Scalars and vectors, Position and Displacement Vectors, Equality of Vectors, Multiplication of vectors by real numbers, Addition and subtraction of vectors - graphical method, Resolution of vectors, Vector addition - analytical method, Motion in a plane, Position Vector and Displacement, Velocity, Acceleration, Motion in a plane with constant acceleration, Relative velocity in two dimensions, Projectile motion, Equation of path of a projectile, Time of Maximum height, Maximum height of a projectile, Horizontal range of projectile, Uniform circular motion.

5) LAWS OF MOTION: Introduction, Aristotle's fallacy, The law of inertia, Newton's first law of motion, Newton's second law of motion, momentum, Impulse, Newton's third law of motion, Conservation of momentum, Equilibrium of a particle, Common forces in mechanics, friction, Circular motion, Motion of a car on a level road, Motion of a car on a Banked road, Solving problems in mechanics.

6) WORK, ENERGY AND POWER: Introduction, The Scalar Product, Notions of work and kinetic energy : The work-energy theorem, Work, Kinetic energy, Work done by a variable force, The work-energy theorem for a variable force, The concept of Potential Energy, The conservation of Mechanical Energy, The Potential Energy of a spring, Various forms of energy: the law of conservation of energy, Heat, Chemical Energy, Electrical Energy, The Equivalence of Mass and Energy, Nuclear Energy, The Principle of Conservation of Energy, Power, Collisions, Elastic and Inelastic Collisions, Collisions in one dimension, Coefficient of Restitution and its determination, Collisions in Two Dimensions.

7) SYSTEMS OF PARTICLES AND ROTATIONAL MOTION: Introduction, What kind of motion can a rigid body have?, Centre of mass, Centre of Gravity, Motion of centre of mass, Linear momentum of a system of particles, Vector product of two vectors, Angular velocity and its relation with linear velocity, Angular acceleration, Kinematics of rotational motion

about a fixed axis, Torque and angular momentum, Moment of force (Torque), Angular momentum of particle, Torque and angular momentum for a system of particles, conservation of angular momentum, Equilibrium of a rigid body, Principle of moments, Moment of inertia, Theorems of perpendicular and parallel axes, Theorem of perpendicular axes, Theorem of parallel axes, Dynamics of rotational motion about a fixed axis, Angular momentum in case of rotation about a fixed axis, Conservation of Angular Momentum, Rolling motion, Kinetic Energy of Rolling Motion.

8) OSCILLATIONS: Introduction, Periodic and oscillatory motions, Period and frequency, Displacement, Simple harmonic motion (S.H.M.), Simple harmonic motion and uniform circular motion, Velocity and acceleration in simple harmonic motion, Force law for Simple harmonic Motion, Energy in simple harmonic motion, some systems executing Simple Harmonic Motion, Oscillations due to a spring, The Simple Pendulum, Damped simple harmonic motion, Forced oscillations and resonance.

9) GRAVITATION: Introduction, Kepler's laws, Universal law of gravitation, The gravitational constant, Acceleration due to gravity of the earth, Acceleration due to gravity below and above the surface of earth, Gravitational potential energy, Escape speed, Earth satellite, Energy of an orbiting satellite, Geostationary and polar satellites, Weightlessness.

10) MECHANICAL PROPERTIES OF SOLIDS: Introduction, Elastic behaviour of solids, Stress and strain, Hooke's law, Stress-strain curve, Elastic moduli, Young's Modulus, Determination of Young's Modulus of the Material of a Wire, Shear Modulus, Bulk Modulus, Poisson's Ratio, Applications of elastic behaviour of materials.

11) MECHANICAL PROPERTIES OF FLUIDS: Introduction, Pressure, Pascal's Law, Variation of Pressure with Depth, Atmosphere Pressure and Gauge Pressure, Hydraulic Machines, Streamline flow, Bernoulli's principle, Speed of Efflux: Torricelli's Law, Venturimeter, Blood Flow and Heart Attack, Dynamic Lift, Viscosity, Variation of Viscosity of fluids with temperature, Stokes' Law, Reynolds number, Surface tension, Surface Energy, Surface Energy and Surface Tension, Angle of Contact, Drops and Bubbles, Capillary Rise, Detergents and Surface Tension.

12) THERMAL PROPERTIES OF MATTER: Introduction, Temperature and heat, Measurement of temperature, Ideal-gas equation and absolute temperature, Thermal expansion, Specific heat capacity, Calorimetry, Change of state, Regelation, Latent Heat, Heat transfer, Conduction, thermal conductivity, Convection, Radiation, Black body Radiation, Greenhouse Effect, Newton's law of cooling.

13) THERMODYNAMICS: Introduction, Thermal equilibrium, Zeroth law of thermodynamics, Heat, Internal Energy and work, First law of thermodynamics, Specific heat capacity, Thermodynamic state variables and equation of State, Thermodynamic process, Quasi-static process, Isothermal Process, Adiabatic Process, Isochoric Process, Isobaric process, Cyclic process, Heat engines, Refrigerators and heat pumps, Second law of thermodynamics, Reversible and irreversible processes, Carnot engine, Carnot's theorem.

14) KINETIC THEORY: Introduction, Molecular nature of matter, Behaviour of gases, Boyle's Law, Charles' Law, Kinetic theory of an ideal gas, Pressure of an Ideal Gas, Law of equipartition of energy, Specific heat capacity, Monatomic Gases, Diatomic Gases, Polyatomic Gases, Specific Heat Capacity of Solids, Specific Heat Capacity of Water, Mean free path.

15) WAVES: Introduction, Transverse and longitudinal waves, Displacement relation in a progressive wave, The speed of a travelling wave, The principle of superposition of waves, Reflection of waves, Beats.

16) RAY OPTICS AND OPTICAL INSTRUMENTS: Introduction, Refraction, Total Internal Reflection, Refraction at Spherical Surfaces and by Lenses, Refraction through a Prism, Dispersion by a Prism, Some Natural Phenomena due to Sunlight, Optical Instruments.

17) WAVE OPTICS: Introduction, Huygens Principle, Refraction and reflection of plane waves using Huygens Principle, Coherent and Incoherent Addition of Waves, Interference of Light Waves and Young's Experiment, Diffraction - The single slit, seeing the single slit diffraction pattern. The validity of ray optics.

18) ELECTRIC CHARGES AND FIELDS: Introduction, Electric Charges, Conductors and Insulators, Charging by Induction, Basic Properties of Electric Charge, Coulomb's Law, Forces between Multiple Charges, Electric Field, Electric Field Lines, Electric Flux, Electric Dipole, Dipole in a Uniform External Field, Continuous Charge Distribution, Gauss's Law, Application of Gauss's Law - Field due to an infinitely long straight uniformly charged wire, Field due to a uniformly charged infinite plane sheet.

19) ELECTROSTATIC POTENTIAL AND CAPACITANCE: Introduction, Electrostatic Potential, Potential due to a Point Charge, Potential due to an Electric Dipole, Potential due to a System of Charges, Equipotential Surfaces, Potential Energy of a System of Charges, Potential Energy in an External Field, Electrostatics of Conductors, Dielectrics and Polarisation, Capacitors and Capacitance, The Parallel Plate Capacitor, Effect of Dielectric on Capacitance, Combination of Capacitors, Energy Stored in a Capacitor.

20) CURRENT ELECTRICITY: Introduction, Electric Current, Electric Currents in Conductors, Ohm's law, Drift of Electrons and the Origin of Resistivity, Limitations of Ohm's Law, Temperature Dependence of Resistivity, Electrical Energy, Power, Cells, emf, Internal Resistance, Cells in Series and in Parallel, Kirchhoff's rules, Wheatstone Bridge, Meter Bridge, Potentiometer.

21) MOVING CHARGES AND MAGNETISM: Introduction, Magnetic Force, Motion in a Magnetic Field, Magnetic Field due to a Current Element, Biot-Savart Law, Magnetic Field on the Axis of a Circular Current Loop, Ampere's Circuital Law, The Solenoid and the Toroid, Force between Two Parallel Currents, the Ampere, Torque on Current Loop, Magnetic Dipole, The Moving Coil Galvanometer.

22) MAGNETISM AND MATTER: Introduction, The Bar Magnet - The magnetic field lines, The Electrostatic analog, Magnetism and Gauss's Law, The Earth's Magnetism, Magnetisation and Magnetic Intensity.

23) ELECTROMAGNETIC INDUCTION: Introduction, The Experiments of Faraday and Henry, Magnetic Flux, Faraday's Law of Induction, Lenz's Law and Conservation of Energy, Motional Electromotive Force, Energy Consideration: A Quantitative Study, Eddy Currents, Inductance, AC Generator.

24) ALTERNATING CURRENT: Introduction, AC Voltage Applied to a Resistor, Representation of AC Current and Voltage by Rotating Vectors-Phasors, AC Voltage Applied

to an Inductor, AC Voltage Applied to a Capacitor, AC Voltage Applied to a Series LCR Circuit, LC Oscillations, Transformers.

25) ELECTROMAGNETIC WAVES: Introduction, Electromagnetic Waves, Electromagnetic Spectrum.

26) DUAL NATURE OF RADIATION AND MATTER: Introduction, Electron Emission, Photoelectric Effect, Experimental Study of Photoelectric Effect, Photoelectric Effect and Wave Theory of Light, Einstein's Photoelectric Equation: Energy Quantum of Radiation, Particle Nature of Light: The Photon, Wave Nature of Matter.

27) ATOMS: Introduction, Alpha-particle Scattering and Rutherford's Nuclear Model of Atom, Atomic Spectra, Bohr Model of the Hydrogen Atom, The Line Spectra of the Hydrogen Atom, DE Broglie's Explanation of Bohr's Second Postulate of Quantisation.

28) NUCLEI: Introduction, Atomic Masses and Composition of Nucleus, Size of the Nucleus, Mass-Energy, Nuclear Force, Nuclear Energy.

29) SEMICONDUCTOR ELECTRONICS: MATERIALS, DEVICES AND SIMPLE CIRCUITS: Introduction, Classification of Materials: Metals, Semiconductors and Insulators, Intrinsic Semiconductor, Extrinsic Semiconductor, p-n Junction, Semiconductor diode, Application of Junction Diode as a Rectifier, Special Purpose p-n Junction Diodes – optoelectronic junction devices, Junction Transistor – transistor - structure and action, basic transistor circuit configurations and transistor characteristics, Digital Electronics and Logic Gates, Integrated Circuits.

30) COMMUNICATION SYSTEMS: Introduction, Elements of a Communication System, Basic Terminology Used in Electronic Communication Systems, Bandwidth of Signals, Bandwidth of Transmission Medium, Propagation of Electromagnetic Waves, Modulation and its Necessity, Amplitude Modulation, Production of Amplitude Modulated Wave, Detection of Amplitude Modulated Wave.

The following topics are the Deleted content of PHYSICS (30%) in the Intermediate Second Year for the Academic Year 2020-21 by the TSBIE as per Telugu Academy text Book. The same is deleted for the TS EAMCET-2021 Examination.

CHAPTER – 1: WAVES

1.8 Doppler Effect

CHAPTER– 2: RAY OPTICS AND OPTICAL INSTRUMENTS

2.2 Reflection of light by Spherical Mirrors

2.8.2 Scattering of light

CHAPTER – 3: WAVE OPTICS

3.6.3 Resolving power of optical instruments

3.7 Polarization

CHAPTER – 4: ELECTRIC CHARGES AND FIELDS

4.15.3 Field due to a uniformly charged thin spherical shell

CHAPTER– 5: ELECTROSTATIC POTENTIAL AND CAPACITANCE

5.16 Van de Graaf generator

CHAPTER – 6: CURRENT ELECTRICITY

6.7 Resistivity of various Materials

6.10 Combination of resistors-series and parallel

CHAPTER – 7: MOVING CHARGES AND MAGNETISM

7.4 Motion in combined electric and magnetic fields

CHAPTER – 8: MAGNETISM AND MATTER

8.2.2 Bar Magnet as a equivalent solenoid

8.2.3 The dipole in a uniform magnetic field

8.6 Magnetic properties of materials

8.7 Permanent magnets and electromagnets

CHAPTER – 10: ALTERNATING CURRENT

10.7 Power in AC Circuit: The Power Factor

CHAPTER – 11: ELECTRO MAGNETIC WAVES

11.2 Displacement Current

CHAPTER–12: DUAL NATURE OF RADIATION AND MATTER

12.9 Davisson and Germer Experiment

CHAPTER–14: NUCLEI

14.4.2 Nuclear Binding energy

14.6 Radioactivity

CHAPTER–15:

SEMICONDUCTOR ELECTRONICS: MATERIALS, DEVICES AND SIMPLE CIRCUITS

15.8.1 Zener diode

15.9.3 Transistor as a device

15.9.4 Transistor as an amplifier (CE configuration)

15.9.5 Feedback amplifier and transistor oscillator

CHEMISTRY

1) ATOMIC STRUCTURE: Introduction; Sub-atomic particles; Atomic models – Thomson's Model; Rutherford's Nuclear model of atom, Drawbacks; Developments to the Bohr's model of atom; Nature of electromagnetic radiation; Particle nature of electromagnetic radiation- Planck's quantum theory; Bohr's model for Hydrogen atom; Explanation of line spectrum of hydrogen; Limitations of Bohr's model; Quantum mechanical considerations of sub atomic particles; Dual behaviour of matter; Heisenberg's uncertainty principle; Quantum mechanical model of an atom. Important features of Quantum mechanical model of atom; Orbitals and quantum numbers; Shapes of atomic orbitals; Energies of orbitals; Filling of orbitals in atoms. Aufbau Principle, Pauli's exclusion Principle and Hund's rule of maximum multiplicity; Electronic configurations of atoms; Stability of half filled and completely filled orbitals.

2) CLASSIFICATION OF ELEMENTS AND PERIODICITY IN PROPERTIES: Need to classify elements; Genesis of periodic classification; Modern periodic law and present form of the periodic table; Nomenclature of elements with atomic number greater than 100; Electronic configuration of elements and the periodic table; Electronic configuration and types of elements s, p, d and f blocks; Trends in physical properties: (a) Atomic radius, (b) Ionic radius (c) Variation of size in inner transition elements, (d) Ionization enthalpy, (e) Electron gain enthalpy, (f) Electro negativity; Periodic trends in chemical properties: (a) Valence or Oxidation states, (b) Anomalous properties of second period elements - diagonal relationship; Periodic trends and chemical reactivity.

3) CHEMICAL BONDING AND MOLECULAR STRUCTURE: Kossel - Lewis approach to chemical bonding, Octet rule, Representation of simple molecules, formal charges, limitations of octet rule; Ionic or electrovalent bond - Factors favourable for the formation of ionic compounds-Crystal structure of sodium chloride, Lattice enthalpy; General properties of ionic compounds; Bond Parameters - bond length, bond angle, bond enthalpy, bond order, resonance - Polarity of bonds dipole moment; Valence Shell Electron Pair Repulsion (VSEPR) theory; Predicting the geometry of simple molecules; Valence bond theory-Orbital overlap concept-Directional properties of bonds-overlapping of atomic orbitals, strength of sigma and pi bonds-Factors favouring the formation of covalent bonds; Hybridisation- different types of hybridization involving s, p and d orbitals- shapes of simple covalent molecules; Coordinate bond -definition with examples; Molecular orbital theory - Formation of molecular orbitals, Linear combination of atomic orbitals (LCAO)-conditions for combination of atomic orbitals - Energy level diagrams for molecular orbitals -Bonding in some homo nuclear diatomic molecules- H_2 , He_2 , Li_2 , B_2 , C_2 , N_2 and O_2 ; Hydrogen bonding-cause of formation of hydrogen bond - Types of hydrogen bonds-inter and intra molecular-General properties of hydrogen bonds.

4) STATES OF MATTER: GASES AND LIQUIDS: Intermolecular forces; Thermal Energy; Intermolecular forces vs Thermal interactions; The Gaseous State; The Gas Laws; Ideal gas equation; Graham's law of diffusion - Dalton's Law of partial pressures; Kinetic molecular theory of gases; Kinetic gas equation of an ideal gas (No derivation); deduction of gas laws from Kinetic gas equation; Distribution of molecular speeds - rms, average and most probable speeds-Kinetic energy of gas molecules; Behaviour of real gases - Deviation from Ideal gas behaviour - Compressibility factor vs Pressure diagrams of real gases; Liquefaction of gases; Liquid State - Properties of Liquids in terms of Inter molecular interactions - Vapour pressure, Viscosity and Surface tension (Qualitative idea only. No mathematical derivation).

5) STOICHIOMETRY: Some basic concepts - Properties of matter - uncertainty in Measurement-significant figures, dimensional analysis; Laws of Chemical Combinations - Law of Conservation of Mass, Law of Definite Proportions, Law of Multiple Proportions, Gay Lussac's Law of Gaseous Volumes, Dalton's Atomic Theory, Avogadro's Law, Atomic and molecular masses- mole concept and molar mass. Concept of equivalent weight; Percentage composition of compounds and calculations of empirical and molecular formulae of compounds; Stoichiometry and stoichiometric calculations; Methods of expressing concentrations of solutions-mass percent, mole fraction, molarity, molality and normality; Redox reactions-classical idea of redox reactions, oxidation and reduction reactions-redox reactions in terms of electron transfer; Oxidation number concept; Types of Redox reactions-combination, decomposition, displacement and disproportionation reactions; Balancing of redox reactions - oxidation number method Half reaction (ion-electron) method; Redox reactions in Titrimetry.

6) THERMODYNAMICS: Thermodynamic Terms; The system and the surroundings; Types of systems and surroundings; The state of the system; The Internal Energy as a State Function. (a) Work (b) Heat (c) The general case, the first law of Thermodynamics; Applications; Work; Enthalpy, H - a useful new state function; Extensive and intensive properties; Heat capacity; The relationship between C_p and C_v ; Measurement of ΔU and ΔH : Calorimetry; Enthalpy change, $\Delta_r H$ of reactions - reaction Enthalpy (a) Standard enthalpy of reactions, (b) Enthalpy changes during phase transformations, (c) Standard enthalpy of formation, (d) Thermo chemical equations (e) Hess's law of constant heat summation; Enthalpies for different types of reactions. (a) Standard enthalpy of combustion ($\Delta_c H^\theta$), (b) Enthalpy of atomization ($\Delta_a H^\theta$), phase transition, sublimation and ionization, (c) Bond Enthalpy ($\Delta_{\text{bond}} H^\theta$), (d) Enthalpy of solution ($\Delta_{\text{sol}} H^\theta$) and dilution; Spontaneity. (a) Is decrease in enthalpy a criterion for spontaneity? (b) Entropy and spontaneity, the second law of thermodynamics, (c) Gibbs Energy and spontaneity; Gibbs Energy change and equilibrium; Absolute entropy and the third law of thermodynamics.

7) CHEMICAL EQUILIBRIUM AND ACIDS-BASES: Equilibrium in Physical processes; Equilibrium in chemical processes - Dynamic Equilibrium; Law of chemical Equilibrium - Law of mass action and Equilibrium constant; Homogeneous Equilibria, Equilibrium constant in gaseous systems. Relationship between K_p and K_c ; Heterogeneous Equilibria; Applications of Equilibrium constant; Relationship between Equilibrium constant K , reaction quotient Q and Gibbs energy G ; Factors affecting Equilibria.-Le Chatelier's principle application to industrial synthesis of Ammonia and Sulphur trioxide; Ionic Equilibrium in solutions; Acids, bases and salts- Arrhenius, Bronsted-Lowry and Lewis concepts of acids and bases; Ionisation of Acids and Bases -Ionisation constant of water and its ionic product- pH scale-ionisation constants of weak acids-ionisation of weak bases-relation between K_a and K_b -Di and poly basic acids and di and poly acidic Bases-Factors affecting acid strength-Common ion effect in the ionization of acids and bases-Hydrolysis of salts and pH of their solutions; Buffer solutions-designing of buffer solution-Preparation of Acidic buffer; Solubility Equilibria of sparingly soluble salts. Solubility product constant, Common ion effect on solubility of ionic salts.

8) HYDROGEN AND ITS COMPOUNDS: Position of hydrogen in the periodic table; Dihydrogen-Occurrence and Isotopes; Preparation and properties of Dihydrogen; Hydrides: Ionic, covalent, and non-stoichiometric hydrides; Water: Physical properties; structure of water, ice. Chemical properties of water; hard and soft water, Temporary and permanent hardness of water; Hydrogen peroxide: Preparation; Physical properties; structure and chemical properties; storage and uses; Heavy Water; Hydrogen as a fuel.

9) THE s - BLOCK ELEMENTS (ALKALI AND ALKALINE EARTH METALS)

Group 1 Elements : Alkali metals; Electronic configurations; Atomic and Ionic radii; Ionization enthalpy; Hydration enthalpy; Physical properties; Chemical properties; Uses; General characteristics of the compounds of the alkali metals: Oxides; Halides; Salts of oxo Acids; Anomalous properties of Lithium: Differences and similarities with other alkali metals, Diagonal relationship; similarities between Lithium and Magnesium; Some important compounds of Sodium: Sodium Carbonate; Sodium Chloride; Sodium Hydroxide; Sodium hydrogen carbonate; Biological importance of Sodium and Potassium.

Group 2 Elements: Alkaline earth elements; Electronic configuration; Ionization enthalpy; Hydration enthalpy; Physical properties, Chemical properties; Uses; General characteristics of compounds of the Alkaline Earth Metals: Oxides, hydroxides, halides, salts of oxoacids (Carbonates; Sulphates and Nitrates); Anomalous behavior of Beryllium; its diagonal relationship with Aluminium; Some important compounds of calcium: Preparation and uses of Calcium Oxide; Calcium Hydroxide; Calcium Carbonate; Plaster of Paris; Cement; Biological importance of Calcium and Magnesium.

10) p- BLOCK ELEMENTS GROUP 13 (BORON FAMILY):

General introduction - Electronic configuration, Atomic radii, Ionization enthalpy, Electro negativity; Physical & Chemical properties; Important trends and anomalous properties of boron; Some important compounds of boron - Borax, Ortho boric acid, diborane; Uses of boron, aluminium and their compounds.

11) p-BLOCK ELEMENTS - GROUP 14 (CARBON FAMILY):

General introduction - Electronic configuration, Atomic radii, Ionization enthalpy, Electro negativity; Physical & Chemical properties; Important trends and anomalous properties of carbon; Allotropes of carbon; Uses of carbon; Some important compounds of carbon and silicon – carbon monoxide, carbon dioxide, Silica, silicones, silicates and zeolites.

12) ENVIRONMENTAL CHEMISTRY:

Definition of terms: Air, Water and Soil Pollutions; Environmental Pollution; Atmospheric pollution; Tropospheric Pollution; Gaseous Air Pollutants (Oxides of Sulphur; Oxides of Nitrogen; Hydrocarbons; Oxides of Carbon (CO, CO₂)). Global warming and Greenhouse effect; Acid rain- Particulate Pollutants- Smog; Stratospheric Pollution: Formation and breakdown of Ozone- Ozone hole- effects of depletion of the Ozone Layer; Water Pollution: Causes of Water Pollution; International standards for drinking water; Soil Pollution: Pesticides, Industrial Wastes; Strategies to control environmental pollution- waste Management- collection and disposal; Green Chemistry: Green chemistry in day-to-day life; Dry cleaning of clothes; Bleaching of paper; Synthesis of chemicals

13) ORGANIC CHEMISTRY-SOME BASIC PRINCIPLES, TECHNIQUES AND HYDROCARBONS

SOME BASIC PRINCIPLES AND TECHNIQUES

General introduction; Tetravalency of Carbon: shapes of organic compounds; Structural representations of organic compounds; Classification of organic compounds; Nomenclature of organic compounds; Isomerism; Fundamental concepts in organic reaction mechanisms; Fission of covalent bond; Nucleophiles and electrophiles; Electron movements in organic reactions; Electron displacement effects in covalent bonds: inductive effect, resonance, resonance effect, electromeric effect, hyperconjugation; Types of Organic Reactions; Methods of purification of organic compounds; Qualitative elemental analysis of organic compounds; Quantitative elemental analysis of organic compounds.

HYDROCARBONS

Classification of Hydrocarbons; **Alkanes** - Nomenclature, isomerism (structural and conformations of ethane only); Preparation of alkanes; Properties - Physical properties and chemical Reactivity, Substitution reactions - Halogenation (free radical mechanism), Combustion, Controlled Oxidation, Isomerisation, Aromatization, reaction with steam and Pyrolysis; **Alkenes** - Nomenclature, structure of ethene, Isomerism (structural and geometrical); Methods of preparation; Properties- Physical and chemical reactions: Addition of Hydrogen, halogen, water, sulphuric acid, Hydrogen halides (Mechanism- ionic and peroxide effect, Markovnikov's, anti-Markovnikov's or Kharasch effect). Oxidation, Ozonolysis and Polymerization; **Alkynes** - Nomenclature and isomerism, structure of acetylene. Methods of preparation of acetylene; Physical properties, Chemical reactions- acidic character of acetylene, addition reactions- of hydrogen, Halogen, Hydrogen halides and water. Polymerization; **Aromatic Hydrocarbons:** Nomenclature and isomerism, Structure of benzene, Resonance and aromaticity; Preparation of benzene. Physical properties. Chemical properties: Mechanism of electrophilic substitution. Electrophilic substitution reactions- Nitration, Sulphonation, Halogenation, Friedel-Crafts alkylation and acylation; Directive influence of functional groups in mono substituted benzene, Carcinogenicity and toxicity

14) SOLID STATE:

General characteristics of solid state; Amorphous and crystalline solids; Classification of crystalline solids based on different binding forces (molecular, ionic, metallic and covalent solids); Probing the structure of solids: X-ray crystallography; Crystal lattices and unit cells. Number of atoms in a unit cell (primitive, body centred and face centred cubic unit cell); Close packed structures: Close packing in one dimension, in two dimensions and in three dimensions- tetrahedral and octahedral voids- formula of a compound and number of voids filled-locating tetrahedral and octahedral voids; Packing efficiency in simple cubic, bcc and in hcp, ccp lattice; Calculations involving unit cell dimensions-density of the unit cell; Imperfections in solids-types of point defects-stoichiometric and non-stoichiometric defects.

15) SOLUTIONS:

Types of solutions; Expressing concentration of solutions - mass percentage, volume percentage, mass by volume percentage, parts per million, mole fraction, molarity and molality; Solubility: Solubility of a solid in a liquid, solubility of a gas in a liquid, Henry's law; Vapour pressure of liquid solutions: vapour pressure of liquid- liquid solutions. Raoult's law as a special case of Henry's law -vapour pressure of solutions of solids in liquids; Ideal and non-ideal solutions; Colligative properties and determination of molar mass-relative lowering of vapour pressure-elevation of boiling point-depression of freezing point-osmosis and osmotic pressure-reverse osmosis and water purification.

16) ELECTROCHEMISTRY AND CHEMICAL KINETICS:

ELECTROCHEMISTRY:

Nernst equation-equilibrium constant from Nernst equation - electro chemical cell and Gibbs energy of the cell reaction; Conductance of electrolytic solutions - measurement of the conductivity of ionic solutions-variation of conductivity and molar conductivity with concentration-strong electrolytes and weak electrolytes-applications of Kohlrausch's law; Electrolysis: Faraday's laws of electrolysis-products of electrolysis.

CHEMICAL KINETICS: Rate of a chemical reaction; Factors influencing rate of a reaction: dependance of rate on concentration- rate expression and rate constant- order of a reaction,

molecularity of a reaction; Integrated rate equations-zero order reactions-first order reactions - half life of a reaction; Pseudo first order reaction; Temperature dependence of the rate of a reaction - Effect of catalyst.

17) SURFACE CHEMISTRY: Adsorption and absorption: Distinction between adsorption and absorption-mechanism of adsorption-types of adsorption-characteristics of physisorption-characteristics of chemisorption - adsorption isotherms-adsorption from solution phase-applications of adsorption.

Colloids; Classification of colloids: Classification based on physical state of dispersed phase and dispersion medium- classification based on nature of interaction between dispersed phase and dispersion medium- classification based on type of particles of the dispersed phase- multi molecular, macromolecular and associated colloids- cleansing action of soaps-preparation of colloids-purification of colloidal solutions- properties of colloidal solutions: Tyndal effect, colour, Brownian movement-charge on colloidal particles, electrophoresis. Coagulation or precipitation, Coagulation of lyophilic sols, protection of colloids, Colloids around us, Applications of colloids.

18) p-BLOCK ELEMENTS:

GROUP-15 ELEMENTS: Occurrence- electronic configuration, atomic and ionic radii, ionisation enthalpy, electronegativity, physical and chemical properties; Dinitrogen- preparation, properties and uses; Compounds of nitrogen-preparation, properties and uses of ammonia; Oxides of nitrogen; Preparation and properties and uses of nitric acid; Phosphorous-allotropic forms;

GROUP-16 ELEMENTS: Occurrence- electronic configuration, atomic and ionic radii, ionisation enthalpy, electron gain enthalpy, electronegativity, physical and chemical properties; Dioxygen-preparation, properties and uses; Simple oxides; Ozone-preparation, properties, structure and uses; Sulphur-allotropic forms; Sulphur dioxide-preparation, properties and uses; Oxoacids of sulphur; Sulphuric acid- properties and uses only.

GROUP-17 ELEMENTS: Occurrence, electronic configuration, atomic and ionic radii, ionisation enthalpy, electron gain enthalpy, electronegativity, physical and chemical properties; Chlorine- preparation, properties and uses; Hydrogen chloride- preparation, properties and uses; Oxoacids of halogens; Interhalogen compounds.

GROUP-18 ELEMENTS: Occurrence, electronic configuration, ionization enthalpy, atomic radii, electron gain enthalpy, physical and chemical properties(a) Xenon-fluorine compounds - XeF_2 , XeF_4 and XeF_6 -preparation, hydrolysis and formation of fluoro anions-structures of XeF_2 , XeF_4 and XeF_6 (b) Xenon-oxygen compounds XeO_3 and XeOF_4 - their formation and structures. Uses.

19) d AND f BLOCK ELEMENTS & COORDINATION COMPOUNDS:

d AND f BLOCK ELEMENTS :Position in the periodic table; Electronic configuration of the d-block elements; General properties of the transition elements (d-block) -physical properties, variation in atomic and ionic sizes of transition series, ionisation enthalpies, oxidation states, trends in the M^{2+}/M and $\text{M}^{3+}/\text{M}^{2+}$ standard electrode potentials, trends in stability of higher oxidation states, chemical reactivity and E^0 values, magnetic properties, formation of coloured ions, formation of complex compounds, catalytic properties, formation of interstitial compounds, alloy formation.

COORDINATION COMPOUNDS: Werner's theory of coordination compounds; Definitions of some terms used in coordination compounds; Nomenclature of coordination compounds-IUPAC nomenclature; Isomerism in coordination compounds-(a)Stereo isomerism-Geometrical and optical isomerism (b)Structural isomerism-linkage,

coordination, ionisation and hydrate isomerism; Bonding in coordination compounds. (a) Valence bond theory - magnetic properties of coordination compounds-limitations of valence bond theory (b) Crystal field theory (i) Crystal field splitting in octahedral and tetrahedral coordination entities (ii) Colour in coordination compounds-limitations of crystal field theory; Bonding in metal carbonyls; Stability of coordination compounds; Importance and applications of coordination compounds.

20) BIOMOLECULES: Carbohydrates- Classification of carbohydrates-Monosaccharides: preparation of glucose from sucrose and starch- Properties and structure of glucose - D, L and (+), (-) configurations of glucose- Structure of fructose.

Amino acids: Natural amino acids - classification of amino acids - structures and D and L forms-Zwitterions; **Proteins:** Structures, classification, fibrous and globular - primary, secondary, tertiary and quaternary structures of proteins- Denaturation of proteins.

Nucleic acids: chemical composition of nucleic acids, structures of nucleic acids, DNA finger printing, biological functions of nucleic acids.

21) HALOALKANES AND HALOARENES: Classification and nomenclature; Nature of C-X bond; Methods of preparation: Alkyl halides and aryl halides-from alcohols, from hydrocarbons (a) by free radical halogenation (b) by electrophilic substitution (c) by replacement of diazonium group (Sandmeyer reaction) (d) by the addition of hydrogen halides and halogens to alkenes-by halogen exchange (Finkelstein reaction); Physical properties-melting and boiling points, density and solubility; Chemical reactions: Reactions of haloalkanes (i) Nucleophilic substitution reactions (a) SN^2 mechanism (b) SN^1 mechanism (c) stereochemical aspects of nucleophilic substitution reactions-optical activity (ii) Elimination reactions (iii) Reaction with metals-Reactions of haloarenes: (i) Nucleophilic substitution (ii) Electrophilic substitution and (iii) Reaction with metals.

22) ORGANIC COMPOUNDS CONTAINING C, H AND O (Alcohols, Phenols, Ethers, Aldehydes, Ketones and Carboxylic acids):

ALCOHOLS, PHENOLS AND ETHERS

Alcohols, phenols and ethers - classification; Nomenclature: (a) Alcohols, (b) phenols and (c) ethers; Structures of hydroxy and ether functional groups; Methods of preparation: **Alcohols** from alkenes and carbonyl compounds (reduction and reaction with Grignard reagents); **Phenols** from haloarenes, benzene sulphonic acid, diazonium salts, cumene; Physical properties of alcohols and phenols; Chemical reactions of alcohols and phenols (i) Reactions involving cleavage of O-H bond-Acidity of alcohols and phenols, esterification (ii) Reactions involving cleavage of C-O bond - reactions with HX, PX_3 , dehydration and oxidation (iii) Reactions of phenols - electrophilic aromatic substitution, Kolbe's reaction, Reimer - Tiemann reaction, reaction with zinc dust, oxidation.

Ethers-Methods of preparation: By dehydration of alcohols, Williamson synthesis - Physical Properties - Chemical reactions: Cleavage of C-O bond and electrophilic substitution of aromatic ethers.

ALDEHYDES AND KETONES

Nomenclature and structure of carbonyl group; Preparation of aldehydes and ketones- (1) by oxidation of alcohols (2) by dehydrogenation of alcohols (3) from hydrocarbons - Preparation of aldehydes (1) from acyl chlorides (2) from nitriles and esters (3) from hydrocarbons-Preparation of ketones (1) from acyl chlorides (2) from nitriles (3) from benzene or substituted benzenes; Physical properties of aldehydes and ketones; Chemical reactions of aldehydes and ketones-nucleophilic addition, reduction, oxidation, reactions due to alpha

Hydrogen and other reactions (Cannizzaro reaction, electrophilic substitution reaction); Uses of aldehydes and ketones.

CARBOXYLIC ACIDS

Nomenclature and structure of carboxyl group; Methods of preparation of carboxylic acids (1) from primary alcohols and aldehydes (2) from alkylbenzenes (3) from nitriles and amides (4) from Grignard reagents (5) from acyl halides and anhydrides (6) from esters; Physical properties; Chemical reactions: (i) Reactions involving cleavage of O-H bond-acidity, reactions with metals and alkalies (ii) Reactions involving cleavage of C-OH bond-formation of anhydride, reactions with PCl_5 , PCl_3 , SOCl_2 , esterification and reaction with ammonia (iii) Reactions involving -COOH group-reduction, decarboxylation (iv) Substitution reactions in the hydrocarbon part - halogenation and ring substitution; Uses of carboxylic acids.

23) ORGANIC COMPOUNDS CONTAINING NITROGEN:

AMINES

Structure of amines; Classification; Nomenclature; Preparation of amines: reduction of nitro compounds, ammonolysis of alkyl halides, reduction of nitriles, reduction of amides, Gabriel phthalimide synthesis and Hoffmann bromamide degradation reaction; Physical properties; Chemical reactions: basic character of amines, alkylation, acylation, carbyl amine reaction, reaction with nitrous acid, reaction with aryl sulphonyl chloride, electrophilic substitution of aromatic amines-bromination, nitration and Sulphonation.

The following topics are the Deleted content of CHEMISTRY (30%) in the Intermediate Second Year for the Academic Year 2020-21 by the TSBIE as per Telugu Academy text Book. The same is deleted for the TS EAMCET-2021 Examination.

Chapter 1 Solid state

1.11 Electrical properties

1.12 Magnetic properties (Band theory of metals, conductors, semiconductors and insulators and n and p type semiconductors)

Chapter 2 Solutions

2.7 Abnormal molar masses-Van't Hoff factor.

Chapter 3 Electrochemistry and Chemical Kinetics

Electrochemistry

3.1 Electrochemical cells

3.2 Galvanic cells

3.5 Electrolytic cells

3.6 Batteries: primary and secondary batteries

3.7 Fuel cells

3.8 Corrosion of metals-Hydrogen economy

Chemical kinetics

3.14 Collision theory of chemical reaction rates.

Chapter 4 Surface Chemistry

4.2 catalysis, homogenous and heterogenous, Adsorption theory of heterogenous catalysis, activity and selectivity of solid catalysis, Shape-selective catalysis by zeolites, Enzyme catalysis, Catalysts in industry.

4.5 Emulsions – types of emulsions

Chapter 5 General Principles of Metallurgy

Entire chapter is deleted

Chapter 6 P-block Elements

Group-15 Elements

6.7 Phosphine-preparation, properties and uses

6.8 Phosphorous halides

6.9 Oxoacids of phosphorous

Group -16 Elements

6.17 Sulphuric acid-manufacture.

Chapter 7 d and f Block Elements & Coordination Compounds

d and f Block Elements

7.4 Some important compounds of transition elements

7.5 Inner transition elements

7.6 Actinoids

7.7 Some applications of d and f block elements.

Chapter 8 Polymers

Entire chapter is deleted

Chapter 9 Biomolecules

9.1 Disaccharides (sucrose, lactose, maltose), polysaccharides (starch, cellulose, glycogen), importance of carbohydrates.

9.3 Enzymes: Enzymes, mechanism of enzyme action

9.4 Vitamins

9.6 Hormones.

Chapter 10 Chemistry in Everyday life

Entire chapter is deleted

Chapter 11 Halo Alkanes and Halo Arenes

11.6 Polyhalogen compounds.

Chapter 12 Organic Compounds Containing C, H and O

12.7 Some Commercially important alcohols (uses with special reference to methanol and ethanol)

Chapter 13 Organic Compounds Containing Nitrogen

II. Diazonium salts

13.7 Methods of preparation of diazonium salts

13.8 Physical properties

13.9 Chemical reactions

13.10 Importance of diazonium salts in synthesis of aromatic compounds

III. Cyanides and Isocyanides

13.11 Structure and nomenclature of cyanides and isocyanides

13.12 Preparation, physical properties and chemical reactions of cyanides and isocyanides.

MODEL QUESTIONS - BOTANY

1. **Assertion (A):** The energy requirement for the formation of one hexose molecule is more
in chlorella than in sugarcane
Reason (R): The formation of phosphoenol pyruvic acid from pyruvic acid requires two ATP molecules.
The correct one is
1) Both (A) and (R) are true and (R) is the correct explanation of (A).
2) Both (A) and (R) are true but (R) is not the correct explanation of (A).
3) (A) is true, (R) is false.
4) (A) is false, (R) is true.
2. Which of the following statements related to protein synthesis are true?
(I) Transcribed hnRNA and mRNA of gene are similar in size.
(II) The 3'CCA and of tRNA acts as amino acid binding site
(III) f-met tRNA binds on to the 'P' site in ribosome.
(IV) The protein factor called RF₃ helps in recognizing the termination codon.
1) (I), (III) 2) (II), (III) 3) (II), (IV) 4) (I), (IV)
3. Inflorescence found in Asteraceae members:
1) Spadix 2) Cyathium 3) Head 4) Umbel
4. Match the following
- | LIST A | LIST B |
|-------------------|-------------|
| A. Monocarpellary | I. Lotus |
| B. Bicarpellary | II. Bean |
| C. Tricarpellary | III. Tomato |
| D. Apocarpous | IV. Onion |
- Identify the correct match between **List-A** and **List-B**
- | | | | | |
|----|----------|----------|---------|----------|
| 1) | A
II | B
IV | C
I | D
III |
| 2) | A
II | B
III | C
IV | D
I |
| 3) | A
III | B
II | C
I | D
IV |
| 4) | A
IV | B
III | C
II | D
I |
5. Naked ovules are the characteristic feature of
1) Bryophytes 2) Angiosperms 3) Gymnosperms 4) Pteridophytes

MODEL QUESTIONS – ZOOLOGY

1. A chronic disorder in which alveolar walls are damaged due to which respiratory surface is decreased. Identify it?
1) Bronchitis 2) Asthma 3) Emphysema 4) Pneumonia
2. Epididymis lined with
1) Stratified cuboidal epithelium 2) Pseudo stratified ciliated epithelium
3) Pseudo stratified non ciliated epithelium 4) Ciliated columnar epithelium
3. Match the following

List-A

- A) Basophils
- B) Acidophils
- C) Neutrophils
- D) Lymphocytes
- E) Monocytes

List-B

- I) Large Cytoplasmic
- II) Peripheral Cytoplasm
- III) Phagocytes
- IV) Small cytoplasmic
- V) Irregular Cytoplasmic

Identify the correct match between **List-A** and **List-B**

1)	A V	B I	C II	D IV	E III
2)	A I	B V	C IV	D II	E III
3)	A V	B I	C III	D IV	E II
4)	A V	B I	C IV	D II	E III

4. Large ecosystem is called
1) Biome 2) Ecotone 3) Community 4) Ecosphere
5. Statement(I)- Red muscle can also be called aerobic muscle.
Statement (II)-These muscle contain plenty of mitochondria which can utilize the large amount of oxygen stored in them for the production of ATP.

1) Both Statement I and Statement II are correct.
2) Both Statement I and Statement II are incorrect.
3) Statement I is correct but Statement II is incorrect
4) Statement I is incorrect but Statement II is correct

MODEL QUESTIONS – PHYSICS

1. A particle performs simple harmonic motion with amplitude **A** and time period **T**. The mean velocity of the particle over the time interval during which it travels a distance of **A/2** starting from extreme position.

1) A/T 2) $2A/T$ 3) $3A/T$ 4) $A/2T$

2. When a battery connected across a resistor of $16\ \Omega$, the voltage across the resistor is 12V. When same battery is connected across a resistor of $10\ \Omega$ voltage across it is 11V. The internal resistance of the battery

1) $10/7\ \Omega$ 2) $20/7\ \Omega$ 3) $25/7\ \Omega$ 4) $30/7\ \Omega$

3. **Assertion (A):** A given force applied in turn to a number of different masses may cause the same rate of change in momentum in each, but not same acceleration to all.

Reason (R): For particles $\mathbf{F} = \frac{dp}{dt}$ and $\mathbf{F} = m\mathbf{a}$ where the symbols have their usual meanings.

- 1) (A) is true & (R) is true and R is the correct explanation of (A)
 2) (A) is true & (R) is true and R is not the correct explanation of (A)
 3) (A) is true & (R) is false
 4) (A) is false & (R) is true

4. **Statement(A):** A particle can have zero displacement and non zero average velocity.

Statement (B): A particle can have zero acceleration and non zero velocity

Statement (C): A particle can have zero velocity and non-zero acceleration.

- 1) A, B, C True 2) A, B True, C False 3) B,C True, A False 4) A,B,C False.

5. Match the following

In the experimental study of photoelectric effect:

Column-I	Column-II
A. Intensity of incident light changes	I. Maximum K.E of photoelectrons changes
B. Frequency of incident light changes	II. Stopping potential changes
C. Target material changes	III. Saturation current changes.

1. A-III, B-I,II C-I,II
 2. A-II, B-I,III C-I,II
 3. A-III, B-III,II C-I,II
 4. A-I, B-I,II C-I,II

MODEL QUESTIONS – CHEMISTRY

- Half-life of a first order reaction is 15min. The time required for completion of 87.5% of the reaction is
 (1) 15 min (2) 30 min (3) 60 min (4) 45 min
- Arrange the following in increasing order of their basic strength.
 I) $\text{C}_2\text{H}_5\text{NH}_2$ II) $\text{C}_6\text{H}_5\text{NH}_2$ III) $\text{C}_6\text{H}_5\text{CH}_2\text{NH}_2$ IV) $(\text{C}_2\text{H}_5)_2\text{NH}$
 (1) $\text{I} < \text{II} < \text{IV} < \text{III}$ (2) $\text{II} < \text{III} < \text{I} < \text{IV}$
 (3) $\text{I} < \text{IV} < \text{III} < \text{II}$ (4) $\text{IV} < \text{I} < \text{III} < \text{II}$
- Which of the following has least electron gain enthalpy?
 (1) Sulphur (2) Chlorine (3) Fluorine (4) Oxygen
- Assertion (A): H_2O is liquid and H_2S is gas at room temperature
 Reason (R) : Molecules of H_2O are highly associated through hydrogen bonding
 The correct answer is:
 (1) Both (A) and (R) are true and (R) is the correct explanation of (A)
 (2) Both (A) and (R) are true and (R) is not the correct explanation of (A)
 (3) (A) is true but (R) is false
 (4) (A) is false but (R) is true

- Match the following:

LIST I (Crystal System)

- (A) Cubic
 (B) Hexagonal
 (C) Monoclinic
 (D) Triclinic

LIST II (Axial Angle)

- (I) $\alpha = \beta = 90^\circ; \gamma = 120^\circ$
 (II) $\alpha \neq \beta \neq \gamma \neq 90^\circ$
 (III) $\alpha = \beta = \gamma \neq 90^\circ$
 (IV) $\alpha = \gamma = 90^\circ; \beta \neq 90^\circ$
 (V) $\alpha = \beta = \gamma = 90^\circ$

The correct match is:

(1)	A	B	C	D
	V	I	III	IV
(2)	A	B	C	D
	IV	II	V	I
(3)	A	B	C	D
	V	I	IV	II
(4)	A	B	C	D
	IV	II	V	III
