
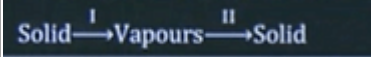
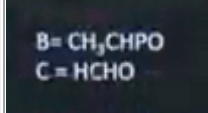
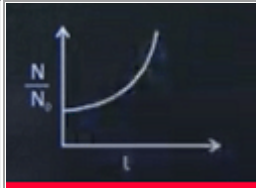
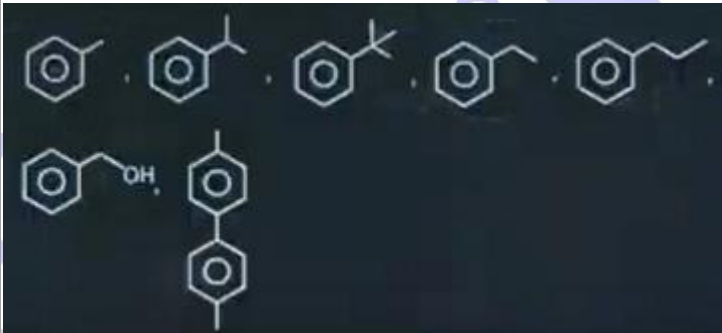


JEE MAIN 28 JANUARY 2025 SHIFT 2

CHEMISTRY QUESTION PAPER WITH ANSWER KEY

Q.No	Questions	Answers										
1	Consider the following oxides: V_2O_3 , V_2O_4 , V_2O_5 Change in oxidation state of vanadium when amphoteric oxide reacts with acids to form VO_4^{3-}	2										
2	Which has maximum oxidising power among the following	MnO_4^-										
3	Number of paramagnetic species among the following is O_2 , O^{2+} , O^{2-} , NO_2 , NO , CO	5										
4	How many of the following molecules are polar? CH_4 , CCl_4 , CH_2Cl_2 , H_2O , NH_3 , H_2O_2 , O_2F_2	5										
5	Which of the following complex is paramagnetic	$\{NiCl_4\}^{2-}$										
6	30 gm of HNO_3 is added to a solution to prepare 75% w/w solution having density 1.25 g/mL. Volume of solution is	32 mL										
7		Both S-I and S-II statements are correct										
8	For an elementary reaction $A + B \rightarrow C + D$ When volume becomes 1/3rd, rate of reaction becomes	9 times										
9	Match the following List-I with List-II. <table border="1" style="margin-left: auto; margin-right: auto;"> <thead> <tr> <th>List-I</th> <th>List-II</th> </tr> </thead> <tbody> <tr> <td>(A) $[COF_6]^{3-}$</td> <td>(i) sp^3d^2</td> </tr> <tr> <td>(B) $[CO(NH_3)_6]^{3+}$</td> <td>(ii) d^2sp^3</td> </tr> <tr> <td>(C) $[NiCl_4]^{2-}$</td> <td>(iii) sp^3</td> </tr> <tr> <td>(D) $[Ni(CN)_4]^{2-}$</td> <td>(iv) dsp^2</td> </tr> </tbody> </table> Choose the correct answer from the options given below:	List-I	List-II	(A) $[COF_6]^{3-}$	(i) sp^3d^2	(B) $[CO(NH_3)_6]^{3+}$	(ii) d^2sp^3	(C) $[NiCl_4]^{2-}$	(iii) sp^3	(D) $[Ni(CN)_4]^{2-}$	(iv) dsp^2	A (i), B (ii), C (iii), D (iv)
List-I	List-II											
(A) $[COF_6]^{3-}$	(i) sp^3d^2											
(B) $[CO(NH_3)_6]^{3+}$	(ii) d^2sp^3											
(C) $[NiCl_4]^{2-}$	(iii) sp^3											
(D) $[Ni(CN)_4]^{2-}$	(iv) dsp^2											
10	The correct name of I & II in the following process is: 	I- Sublimation II-Deposition										
11	Consider the following sequence of reaction $CH_3 - C \equiv CH \xrightarrow[\text{CaCO}_3]{H_2/Pd} A \xrightarrow[H_2O]{O_3/Zn} B + C$											
12	Which of the following biomolecules doesn't contain C_1 - C_4 glycosidic linkage	Sucrose										
13	Consider the following statements	Both Statements I										

	<p>Statement I: In law of octaves, elements were arranged in ascending order of their atomic numbers.</p> <p>Statement II: Lothar Meyer, plotted the physical properties against atomic weight.</p> <p>Choose the correct answer from the options given below.</p>	and II are incorrect
14	<p>Consider the following sequence of reaction</p> $C_6H_{12} \xrightarrow{Se/\Delta} A \xrightarrow[AlCl_3]{CH_3-Cl} B \xrightarrow[H_3O^+]{CrO_2Cl_2} C$ <p>Choose the correct option about major product</p>	'C' can give Tollen's test
15	<p>Consider the following oxides V_2O_3, V_2O_4, V_2O_5</p> <p>Oxidation state of vanadium in amphoteric oxide is</p>	+5
16	<p>The bacterial life grows as per 1st order of kinetics. Which of the following graph is correct between N/N_0 and t?</p>	
17	<p>How many of the following will give Benzoic acid on reaction with hot alkaline $KMnO_4$?</p> 	5
18	<p>By passing current in 600 mL of NaCl solution pH increase to 12.</p> <p>Find current (i) if electrolysis occurs for 10 min/assume 100% efficiency</p>	0.965