

CHEMISTRY

SECTION-A

51. Given below are two statements : one is labelled as **Assertion A** and the other is labelled as **Reason R** :

Assertion A : A reaction can have zero activation energy.

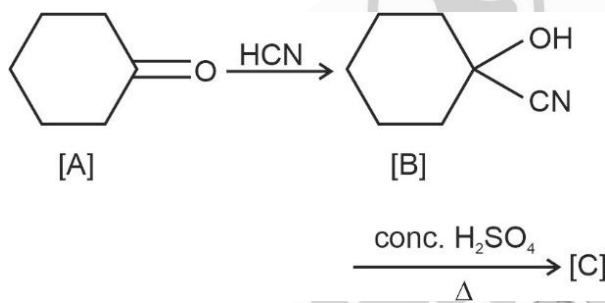
Reasons R : The minimum extra amount of energy absorbed by reactant molecules so that their energy becomes equal to threshold value, is called activation energy.

In the light of the above statements, choose the **correct** answer from the options given below :

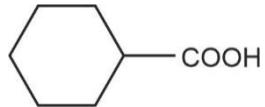
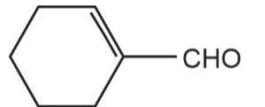
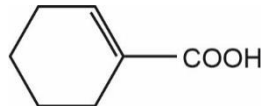
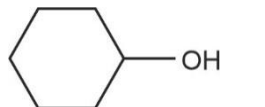
- (1) Both **A** and **R** are true and **R** is **NOT** the correct explanation of **A**
- (2) **A** is true but **R** is false
- (3) **A** is false but **R** is true
- (4) Both **A** and **R** are true and **R** is the correct explanation of **A**

Answer (1)

52. Complete the following reaction

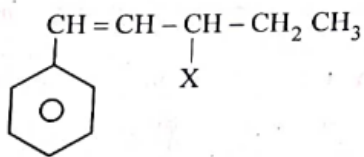


[C] is _____

- (1) 
- (2) 
- (3) 
- (4) 

Answer (3)

53. The given compound



is an example of _____.

- (1) Aryl halide
- (2) Allylic halide
- (3) Vinylic halide
- (4) Benzylic halide

Answer (2)

54. Intermolecular forces are forces of attraction and repulsion between interacting particles that will include :

- A. dipole - dipole forces
- B. dipole - induced dipole forces
- C. hydrogen bonding
- D. covalent bonding
- E. dispersion forces

Choose the **most appropriate** answer from the options given below :

- (1) A, B, C, D are correct
- (2) A, B, C, E are correct
- (3) A, C, D, E are correct
- (4) B, C, D, E are correct

Answer (2)

55. Which of the following reactions will NOT give primary amine as the product?

- (1) $\text{CH}_3\text{CN} \xrightarrow[\text{(ii) H}_3\text{O}^{\oplus}}{\text{(i) LiAlH}_4} \text{Product}$
- (2) $\text{CH}_3\text{NC} \xrightarrow[\text{(ii) H}_3\text{O}^{\oplus}}{\text{(i) LiAlH}_4} \text{Product}$
- (3) $\text{CH}_3\text{CONH}_2 \xrightarrow[\text{(ii) H}_3\text{O}^{\oplus}}{\text{(i) LiAlH}_4} \text{Product}$
- (4) $\text{CH}_3\text{CONH}_2 \xrightarrow{\text{Br}_2/\text{KOH}} \text{Product}$

Answer (2)

56. Given below are two statements : one is labelled as **Assertion A** and the other is labelled as **Reason R** :

Assertion A : Metallic sodium dissolves in liquid ammonia giving a deep blue solution, which is paramagnetic.

Reason R : The deep blue solution is due to the formation of amide.

In the light of the above statements, choose the **correct** answer from the options given below :

- (1) Both **A** and **R** are true but **R** is **NOT** the correct explanation of **A**
- (2) **A** is true but **R** is false
- (3) **A** is false but **R** is true
- (4) Both **A** and **R** are true and **R** is the correct explanation of **A**

Answer (2)

57. For a certain reaction, the rate = $k[A]^2[B]$, when the initial concentration of A is tripled keeping concentration of B constant, the initial rate would

- (1) Increase by a factor of six
- (2) Increase by a factor of nine
- (3) Increase by a factor of three
- (4) Decrease by a factor of nine

Answer (2)

58. Taking stability as the factor, which one of the following represents **correct** relationship?

- (1) $\text{InI}_3 > \text{InI}$
- (2) $\text{AlCl} > \text{AlCl}_3$
- (3) $\text{TlI} > \text{TlI}_3$
- (4) $\text{TlCl}_3 > \text{TlCl}$

Answer (3)

59. Amongst the given options which of the following molecules/ ion acts as a Lewis acid?

- (1) H_2O
- (2) BF_3
- (3) OH^-
- (4) NH_3

Answer (2)

60. The number of σ bonds, π bonds and lone pair of electrons in pyridine, respectively are:

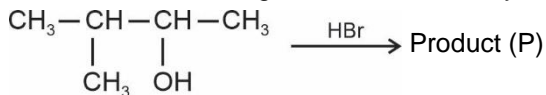
- (1) 12, 3, 0
- (2) 11, 3, 1
- (3) 12, 2, 1
- (4) 11, 2, 0

Answer (2)

61. The **correct** order of energies of molecular orbitals of N_2 molecule, is
- (1) $\sigma 1s < \sigma^* 1s < \sigma 2s < \sigma^* 2s < \sigma 2p_z < (\pi 2p_x = \pi 2p_y) < (\pi^* 2p_x = \pi^* 2p_y) < \sigma^* 2p_z$
 - (2) $\sigma 1s < \sigma^* 1s < \sigma 2s < \sigma^* 2s < \sigma 2p_z < \sigma^* 2p_z < (\pi 2p_x = \pi 2p_y) < (\pi^* 2p_x = \pi^* 2p_y)$
 - (3) $\sigma 1s < \sigma^* 1s < \sigma 2s < \sigma^* 2s < (\pi 2p_x = \pi 2p_y) < (\pi^* 2p_x = \pi^* 2p_y) < \sigma 2p_z < \sigma^* 2p_z$
 - (4) $\sigma 1s < \sigma^* 1s < \sigma 2s < \sigma^* 2s < (\pi 2p_x = \pi 2p_y) < \sigma 2p_z < (\pi^* 2p_x = \pi^* 2p_y) < \sigma^* 2p_z$

Answer (4)

62. Consider the following reaction and identify the product (P).



3-Methylbutan-2-ol

- (1) $\text{CH}_3\text{CH} = \text{CH} - \text{CH}_3$
- (2) $\begin{array}{c} \text{CH}_3 - \text{CH} - \text{CH} - \text{CH}_3 \\ | \quad | \\ \text{CH}_3 \quad \text{Br} \\ | \\ \text{CH}_3 \end{array}$
- (3) $\begin{array}{c} \text{CH}_3 - \text{C} - \text{CH}_2\text{Br} \\ | \\ \text{CH}_3 \\ | \\ \text{Br} \end{array}$
- (4) $\begin{array}{c} \text{CH}_3 - \text{C} - \text{CH}_2 - \text{CH}_3 \\ | \\ \text{CH}_3 \end{array}$

Answer (4)

63. Given below are two statements: one is labelled as **Assertion A** and the other is labelled as **Reason R**

Assertion A : In equation $\Delta_r G = -nFE_{\text{cell}}$ value of $\Delta_r G$ depends on n .

Reasons R : E_{cell} is an intensive property and $\Delta_r G$ is an extensive property.

In the light of the above statements, choose the **correct** answer from the options given below

- (1) Both **A** and **R** are true and **R** is **NOT** the correct explanation of **A**
- (2) **A** is true but **R** is false
- (3) **A** is false but **R** is true
- (4) Both **A** and **R** are true and **R** is the correct explanation of **A**

Answer (1)

64. Homoleptic complex from the following complexes is

- (1) Diamminechloridonitrito-N-platinum (II)
- (2) Pentaamminecarbonatocobalt (III) chloride
- (3) Triamminetriaquachromium (III) chloride
- (4) Potassium trioxalatoaluminate (III)

Answer (4)

65. Given below are two statements : one is labelled as **Assertion A** and the other is labelled as **Reason R**

Assertion A : Helium is used to dilute oxygen in diving apparatus.

Reason R : Helium has high solubility in O_2 .

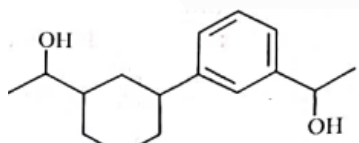
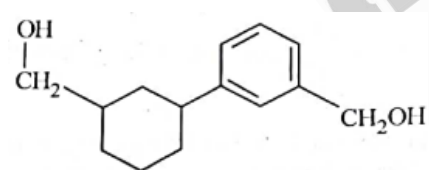
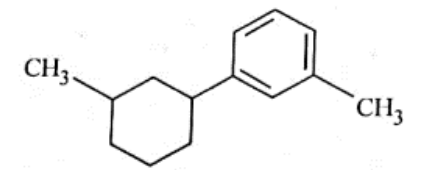
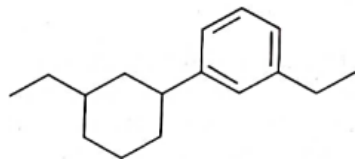
In the light of the above statements, choose the **correct** answer from the options given below

- (1) Both **A** and **R** are true and **R** is **NOT** the correct explanation of **A**
- (2) **A** is true but **R** is false
- (3) **A** is false but **R** is true
- (4) Both **A** and **R** are true and **R** is the correct explanation of **A**

Answer (1)

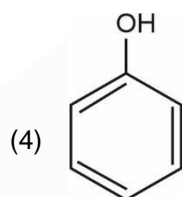
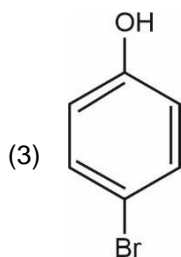
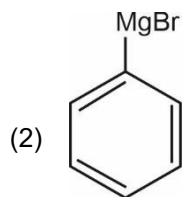
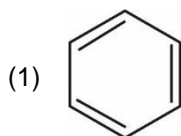
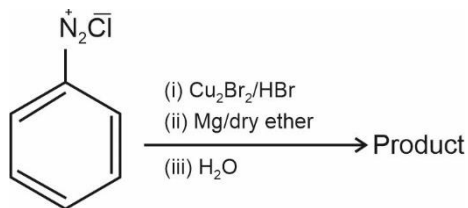
66. Identify product (A) in the following reaction:



- (1) 
- (2) 
- (3) 
- (4) 

Answer (4)

67. Identify the product in the following reaction:



Answer (1)

68. Given below are two statements :

Statement I : A unit formed by the attachment of a base to 1' position of sugar is known as nucleoside.

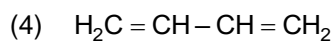
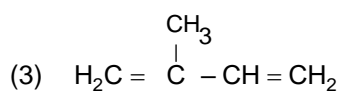
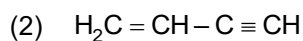
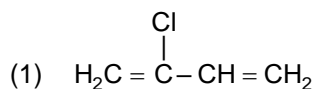
Statement II : When nucleoside is linked to phosphorous acid at 5' -position of sugar moiety, we get nucleotide.

In the light of the above statements, choose the **correct** answer from the options given below :

- (1) Both Statement I and Statement II are false
- (2) Statement I is true but Statement II is false
- (3) Statement I is false but Statement II is true
- (4) Both Statement I and Statement II are true

Answer (2)

69. Which amongst the following molecules on polymerization produces neoprene?



Answer (1)

70. The relation between n_m , (n_m = the number of permissible values of magnetic quantum number (m)) for a given value of azimuthal quantum number (l), is

(1) $l = 2n_m + 1$

(2) $n_m = 2l + 1$

(3) $n_m = l + 2$

(4) $l = \frac{n_m - 1}{2}$

Answer (4)

71. The stability of Cu^{2+} is more than Cu^+ salts in aqueous solution due to

(1) Enthalpy of atomization

(2) Hydration energy

(3) Second ionisation enthalpy

(4) First ionisation enthalpy

Answer (2)

72. Some tranquilizers are listed below. Which one from the following belongs to barbiturates?

(1) Meprobamate

(2) Valium

(3) Veronal

(4) Chlordiazepoxide

Answer (3)

73. The element expected to form largest ion to achieve the nearest noble gas configuration is

(1) F

(2) N

(3) Na

(4) O

Answer (2)

74. Which of the following statements are **NOT** correct?

A. Hydrogen is used to reduce heavy metal oxides to metals.

B. Heavy water is used to study reaction mechanism.

C. Hydrogen is used to make saturated fats from oils.

D. The H–H bond dissociation enthalpy is lowest as compared to a single bond between two atoms of any element.

E. Hydrogen reduces oxides of metals that are more active than iron.

Choose the **most appropriate** answer from the options given below:

(1) B, D only

(2) D, E only

(3) A, B, C only

(4) B, C, D, E only

Answer (2)

75. A compound is formed by two elements A and B. The element B forms cubic close packed structure and atoms of A occupy $\frac{1}{3}$ of tetrahedral voids. If the formula of the compound is A_xB_y , then the value of $x + y$ is in option

- (1) 4 (2) 3
(3) 2 (4) 5

Answer (4)

76. Which one of the following statements is **correct**?

- (1) All enzymes that utilise ATP in phosphate transfer require Ca as the cofactor
(2) The bone in human body is an inert and unchanging substance
(3) Mg plays roles in neuromuscular function and interneuronal transmission
(4) The daily requirement of Mg and Ca in the human body is estimated to be 0.2-0.3 g

Answer (4)

77. Amongst the following the total number of species NOT having eight electrons around central atom in its outermost shell, is

$NH_3, AlCl_3, BeCl_2, CCl_4, PCl_5$:

- (1) 2 (2) 4
(3) 1 (4) 3

Answer (4)

78. Weight (g) of two moles of the organic compound, which is obtained by heating sodium ethanoate with sodium hydroxide in presence of calcium oxide is :

- (1) 32 (2) 30
(3) 18 (4) 16

Answer (1)

79. Select the **correct** statements from the following

- A. Atoms of all elements are composed of two fundamental particles.
B. The mass of the electron is 9.10939×10^{-31} kg.
C. All the isotopes of a given element show same chemical properties.
D. Protons and electrons are collectively known as nucleons.
E. Dalton's atomic theory, regarded the atom as an ultimate particle of matter

Choose the **correct** answer from the options given below

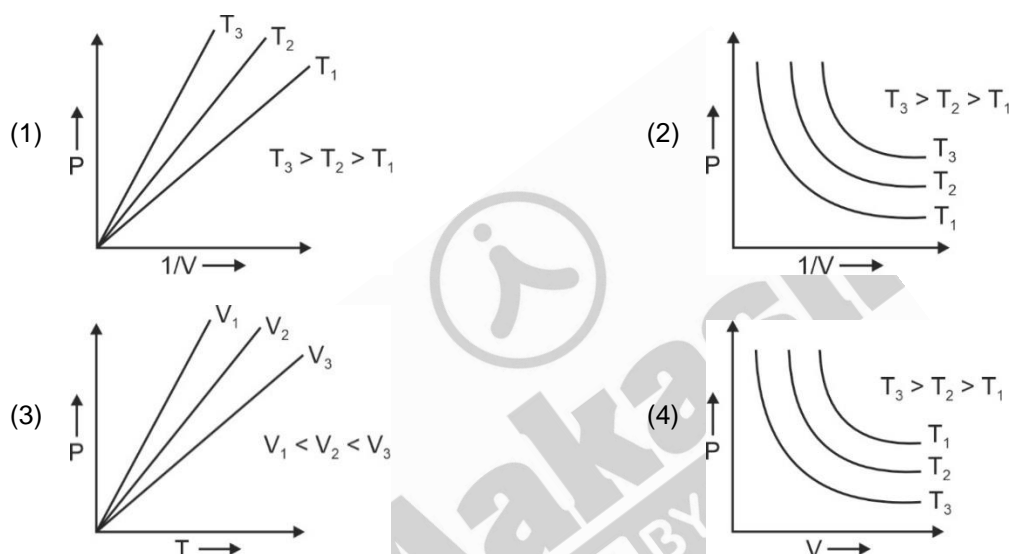
- (1) C, D and E only (2) A and E only
(3) B, C and E only (4) A, B and C only

Answer (3)

80. Which one is an example of heterogenous catalysis?
- (1) Hydrolysis of sugar catalysed by H^+ ions
 - (2) Decomposition of ozone in presence of nitrogen monoxide
 - (3) Combination between dinitrogen and dihydrogen to form ammonia in the presence of finely divided iron
 - (4) Oxidation of sulphur dioxide into sulphur trioxide in the presence of oxides of nitrogen

Answer (3)

81. Which amongst the following options is **correct** graphical representation of Boyle's law?



Answer (1)

82. Match **List-I** with **List-II**.

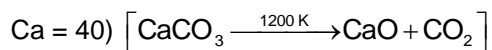
List-I	List-II
A. Coke	I. Carbon atoms are sp^3 hybridised
B. Diamond	II. Used as a dry lubricant
C. Fullerene	III. Used as a reducing agent
D. Graphite	IV. Cage like molecules

Choose the **correct** answer from the options given below :

- (1) A-IV, B-I, C-II, D-III
- (2) A-III, B-I, C-IV, D-II
- (3) A-III, B-IV, C-I, D-II
- (4) A-II, B-IV, C-I, D-III

Answer (2)

83. The **right** option for the mass of CO₂ produced by heating 20 g of 20% pure limestone is (Atomic mass of



- (1) 1.76 g
- (2) 2.64 g
- (3) 1.32 g
- (4) 1.12 g

Answer (1)

84. In Lassaigne's extract of an organic compound, both nitrogen and sulphur are present, which gives blood red colour with Fe³⁺ due to the formation of

- (1) NaSCN
- (2) [Fe(CN)₅NOS]⁴⁻
- (3) [Fe(SCN)]²⁺
- (4) Fe₄[Fe(CN)₆]₃·xH₂O

Answer (3)

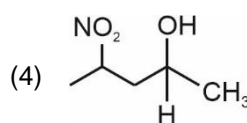
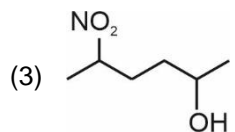
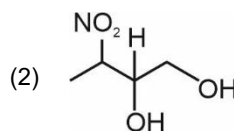
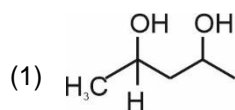
85. The conductivity of centimolar solution of KCl at 25°C is 0.0210 ohm⁻¹ cm⁻¹ and the resistance of the cell containing the solution at 25°C is 60 ohm. The value of cell constant is

- (1) 3.28 cm⁻¹
- (2) 1.26 cm⁻¹
- (3) 3.34 cm⁻¹
- (4) 1.34 cm⁻¹

Answer (2)

SECTION-B

86. Which amongst the following will be most readily dehydrated under acidic conditions?



Answer (1)

87. Match List-I with List-II :

List-I (Oxoacids of Sulphur)

List-II (Bonds)

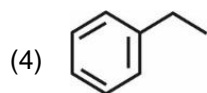
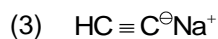
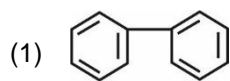
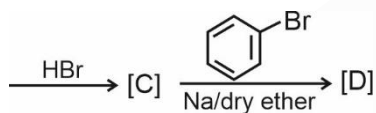
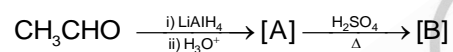
- | | |
|---------------------------|--------------------------------------|
| A. Peroxodisulphuric acid | I. Two S–OH, Four S=O, One S–O–S |
| B. Sulphuric acid | II. Two S–OH, One S=O |
| C. Pyrosulphuric acid | III. Two S–OH, Four S=O, One S–O–O–S |
| D. Sulphurous acid | IV. Two S–OH, Two S=O |

Choose the **correct** answer from the options given below.

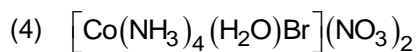
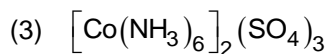
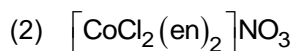
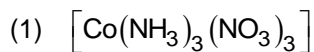
- (1) A–III, B–IV, C–I, D–II
 (2) A–I, B–III, C–IV, D–II
 (3) A–III, B–IV, C–II, D–I
 (4) A–I, B–III, C–II, D–IV

Answer (1)

88. Identify the final product [D] obtained in the following sequence of reactions.

**Answer (4)**

89. Which complex compound is most stable?

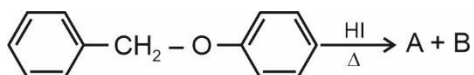
**Answer (2)**

90. Which amongst the following options is the **correct** relation between change in enthalpy and change in internal energy?

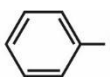
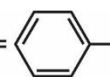
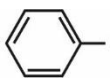
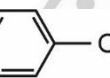
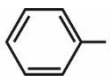
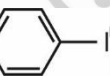
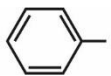
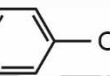
- (1) $\Delta H = \Delta U + \Delta n_g RT$
- (2) $\Delta H - \Delta U = -\Delta n RT$
- (3) $\Delta H + \Delta U = \Delta n R$
- (4) $\Delta H = \Delta U - \Delta n_g RT$

Answer (1)

91. Consider the following reaction :



Identify products A and B.

- (1) A =  and B = 
- (2) A =  and B = 
- (3) A =  and B = 
- (4) A =  and B = 

Answer (2)

92. Given below are two statements :

Statement I : The nutrient deficient water bodies lead to eutrophication

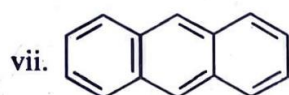
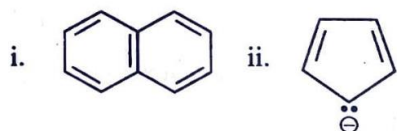
Statement II : Eutrophication leads to decrease in the level of oxygen in the water bodies.

In the light of the above statements, choose the **correct** answer from the options given below:

- (1) Both **Statement I** and **Statement II** are false.
- (2) **Statement I** is correct but **Statement II** is false.
- (3) **Statement I** is incorrect but **Statement II** is true.
- (4) Both **Statement I** and **Statement II** are true.

Answer (3)

93. Consider the following compounds/species:

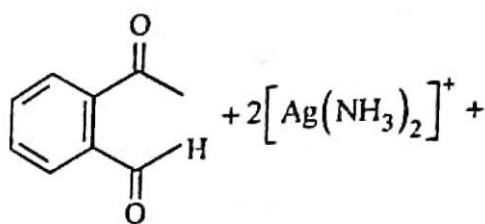


The number of compounds/species which obey Huckel's rule is _____.

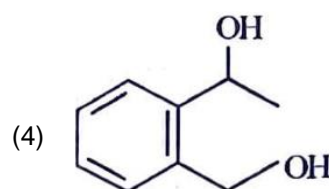
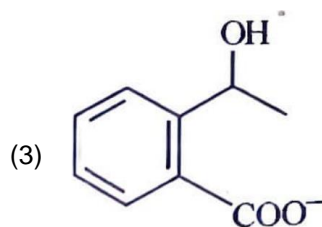
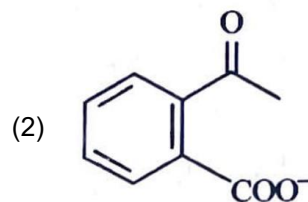
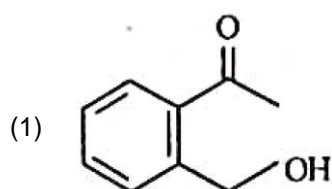
- (1) 6 (2) 2
(3) 5 (4) 4

Answer (4)

94. Identify the major product obtained in the following reaction:



$3^- \text{OH} \xrightarrow{\Delta}$ major product



Answer (2)

95. What fraction of one edge centred octahedral void lies in one unit cell of fcc?

(1) $\frac{1}{3}$

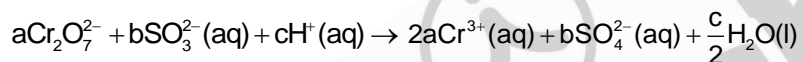
(2) $\frac{1}{4}$

(3) $\frac{1}{12}$

(4) $\frac{1}{2}$

Answer (2)

96. On balancing the given redox reaction,



the coefficients a, b and c are found to be, respectively-

(1) 3, 8, 1

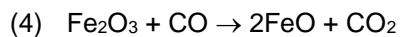
(2) 1, 8, 3

(3) 8, 1, 3

(4) 1, 3, 8

Answer (4)

97. The reaction that does **NOT** take place in a blast furnace between 900 K to 1500 K temperature range during extraction of iron is :



Answer (4)

98. Which of the following statements are **INCORRECT**?
- A. All the transition metals except scandium form MO oxides which are ionic.
 - B. The highest oxidation number corresponding to the group number in transition metal oxides is attained in Sc_2O_3 to Mn_2O_7 .
 - C. Basic character increases from V_2O_3 to V_2O_4 to V_2O_5 .
 - D. V_2O_4 dissolves in acids to give VO_4^{3-} salts.
 - E. CrO is basic but Cr_2O_3 is amphoteric.

Choose the **correct** answer from the options given below:

- (1) B and D only
- (2) C and D only
- (3) B and C only
- (4) A and E only

Answer (2)

99. The equilibrium concentrations of the species in the reaction $\text{A} + \text{B} \rightleftharpoons \text{C} + \text{D}$ are 2, 3, 10 and 6 mol L^{-1} , respectively at 300 K. ΔG° for the reaction is ($R = 2 \text{ cal/mol K}$)
- (1) -137.26 cal
 - (2) -1381.80 cal
 - (3) -13.73 cal
 - (4) 1372.60 cal

Answer (2)

100. Pumice stone is an example of
- (1) Gel
 - (2) Solid sol
 - (3) Foam
 - (4) Sol

Answer (2)