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| :--- | :--- | :--- | :--- | :--- | :--- | :--- |

## PART - III

## CHEMISTRY /

( Urdu \& English Version / اروواورانگريٌكزبان (
[ 3.00 , 3 [
[ 70
Time Allowed : 3.00 Hours ]
[ Maximum Marks : 70

Instructions : (1) Check the question paper for fairness of printing. If there is any lack of fairness, inform the Hall Supervisor immediately.
(2) Use Blue or Black ink to write and underline and pencil to draw diagrams.

Note : Draw diagrams and write equations wherever necessary.

## PART - I/ I-

$15 \times 1=15$

$$
\begin{align*}
& \text { (i) : } \\
& \text {, وَيُ گَش إِتباللات } \tag{ii}
\end{align*}
$$

Note : (i) Answer all the questions.
(ii) Choose the most appropriate answer from the given four alternatives and write the option code and the corresponding answer.



Zr (impure) $+2 \mathrm{I}_{2} \xrightarrow{523 \mathrm{~K}} \mathrm{ZrI}_{4}$ $\mathrm{ZrI}_{4} \xrightarrow{1800 \mathrm{~K}} \mathrm{Zr}$ (pure) $+2 \mathrm{I}_{2}$


し
(b)

(a)

واناركّ
(d)

6انُّ
(c)

The following set of reactions are used in refining Zirconium. This method is called as $\qquad$ .

Zr (impure) $+2 \mathrm{I}_{2} \xrightarrow{523 \mathrm{~K}} \mathrm{ZrI}_{4}$
$\mathrm{ZrI}_{4} \xrightarrow{1800 \mathrm{~K}} \mathrm{Zr}$ (pure) $+2 \mathrm{I}_{2}$
(a) Zone refining
(b) Liquation
(c) Mond's process
(d) Van Arkel process

2
گراینـ
(d)
"
(c) گرافاتُ
(b)
نرلين
(a)

Which of the following is not $\mathrm{sp}^{2}$ hybridised ?
(a) Fullerene
(b) Graphite
(c) Diamond
(d) Graphene

.3
$\mathrm{XeO}_{3}$
(d)
$\mathrm{NeF}_{2}$
(c)
$\mathrm{XeOF}_{4}$
(b)
$\mathrm{XeF}_{2}$
(a)

Which one of the following compounds is not formed ?
(a) $\mathrm{XeF}_{2}$
(b) $\mathrm{XeOF}_{4}$
(c) $\mathrm{NeF}_{2}$
(d) $\mathrm{XeO}_{3}$

al-1-(Butan-1-al)
(b)
(a)
(d)
(c) (
$\mathrm{CH}_{3}-\mathrm{CHO}+\mathrm{CO} \xrightarrow{\mathrm{Rh} / \text { Ir complex }}$ ?
(a) Poly propylene
(b) Butan-1-al
(c) Acetic acid
(d) Acetate

6 (d)
3 (c)
2 (b)
4 (a)

In $\mathrm{K}_{4}\left[\mathrm{Fe}(\mathrm{CN})_{6}\right]$, the co-ordination number of $\mathrm{Fe}^{2+}$ is $\qquad$ .
(a) 4
(b) 2
(c) 3
(d) 6

FeO (d)
KCl
(c)

NaCl
(b)

ZnO
(a)

The crystal with a metal deficiency defect is :
(a) ZnO
(b) NaCl
(c) KCl
(d) FeO
 :


Among the following graphs showing variation of rate constant with temperature ( T ) for a reaction, the one that exhibits Arrhenius behaviour over the entire temperature range is :
(a)

(b)

(c)

(d) Both (b) and (c)

(a)
بك6ا تيزاب (ايمُ)
(b)
UL1
(c)
رزّزتزاب(ايمُ)
(d)

The pH of an aqueous solution is zero. The solution is :
(a) Neutral
(b) Slightly acidic
(c) Basic
(d) Strongly acidic
$6.22 \times 10^{23}$
(b)
$6.022 \times 10^{22}$
(a)
$6.022 \times 10^{24}$
(d)
$6.022 \times 10^{-34}$
(c)

The number of electrons that have a total Charge of 9650 Coulombs is :
(a) $6.022 \times 10^{22}$
(b) $6.22 \times 10^{23}$
(c) $6.022 \times 10^{-34}$
(d) $6.022 \times 10^{24}$


The phenomenon observed when a beam of light is passed through a colloidal solution is :
(a) Coagulation
(b) Cataphoresis
(c) Tyndall effect
(d) Electrophoresis

> 11. جورُّاحيَ:
(i)
(1)
(ii)
(iii)
(2)
(iv)
?
(3)
(1) - (iv), (2) - (iii), (3) - (ii), (4) - (i) (a)

U
(4)
(1) - (iii), (2) - (i), (3) - (ii), (4) - (iv)
(b)
(1) - (ii), (2) - (i), (3) - (iv), (4) - (iii) (c)
(1) - (iii), (2) - (iv), (3) - (i), (4) - (ii)
(d)

Match the following :
(1) Emulsion
(2) Gel
(3) Foam
(4) Sol
(i) Whipped Cream
(ii) Ink
(iii) Cream
(iv) Butter
(a) (1) - (iv), (2) - (iii), (3) - (ii), (4) - (i)
(b) (1) - (iii), (2) - (i), (3) - (ii), (4) - (iv)
(c) (1) - (ii), (2) - (i), (3) - (iv), (4) - (iii)
(d) (1) - (iii), (2) - (iv), (3) - (i), (4) - (ii)
$\qquad$


On reacting with neutral Ferric Chloride, Phenol gives :
(a) Dark green colour
(b) Red colour
(c) No colouration
(d) Violet colour


وكمّميرشُ
(b)
(d)

$$
\begin{aligned}
& \text { كصول 'X' 'نيّن يتابـ } \\
& \text { آيوُوْفا رمنّط } \\
& \text { (a) }
\end{aligned}
$$

$$
\begin{aligned}
& \text { (c) }
\end{aligned}
$$

In the following reaction,
$\mathrm{HC} \equiv \mathrm{CH} \xrightarrow[\mathrm{HgSO}_{4}]{\mathrm{H}_{2} \mathrm{SO}_{4}} \mathrm{X}$
Product ' X ' will not give :
(a) Iodoform test
(b) Tollen's test
(c) Fehling solution test
(d) Victor Meyer test


(b)
(d)

(c)

IUPAC name for the amine $\mathrm{H}_{2} \mathrm{~N}-\mathrm{CH}_{2}-\left(\mathrm{CH}_{2}\right)_{4}-\mathrm{CH}_{2}-\ddot{\mathrm{N}} \mathrm{H}_{2}$
(a) Heptane-1, 7-diamine
(b) Hexamethylene diamine
(c) Hexane-1, 6-amine
(d) Hexane-1, 6-diamine

The pyrimidine bases present in DNA are :
(a) Cytosine and Thiamine
(b) Cytosine and Adenine
(c) Cytosine and Uracil
(d) Cytosine and Guanine

## (إرط- PART - II/ II

## $6 \times 2=12$

$$
\begin{aligned}
& \text { لازنى }
\end{aligned}
$$

Note : Answer any six questions. Question No. 24 is compulsory.

Give the uses of Borax.

$$
17 .
$$

Why 'd' block elements exhibit variable oxidation state ?

Define unit cell.

State Ostwald's dilution law.

Define equivalent conductance.

Mention any two factors that affect electrolytic conductance.
اليكمُورآ آموس س كيامراوب؟

What is meant by Electro Osmosis ?

Write a short note on peptide bond.

From the following reaction, identify A and B.


## PART - III / III (پر

## $6 \times 3=18$

لازگى

نو
Note : Answer any six questions. Question No. 33 is compulsory.

What is meant by term "Coordination Number" ? What is the Coordination Number of atoms in a bcc structure ?

What are interhalogen compounds ? Give two examples.

Give the difference between double salts and coordination compounds.

Mention the factors responsible for the anomalous behaviour of the first element of p-block.

State Faraday's laws of electrolysis.

$$
\begin{align*}
& .30 \\
& \text { (i) } \\
& \text { كأكيمرال } \tag{ii}
\end{align*}
$$

How are the following conversions effected ?
(i) ethylene glycol $\rightarrow$ acetaldehyde
(ii) glycerol $\rightarrow$ acrolein

Give the tests for carboxylic acid group.

Give any three differences between DNA and RNA.


Classify the following into covalent, molecular, ionic and metallic solids.
(i) Diamond
(ii) Brass
(iii) NaCl
(iv) Naphthalene
(v) Glucose
(vi) $\mathrm{SiO}_{2}$

PART - IV / IV إرط

## 5x5=25


Note : Answer all the questions.

(a) Explain Froth flotation process.
(b) (i) Explain the bleaching action of Sulphur dioxide
(ii) Write any two uses of Helium.

(a) (i) What are interstitial compounds ?
(ii) Calculate the number of unpaired electrons in $\mathrm{Ti}^{3+}, \mathrm{Mn}^{2+}$ and calculate the spin only magnetic moment.

## OR

(b) (i) What are the limitations of VB theory ?
(ii) Based on the VB theory, explain why $\left[\mathrm{Ni}(\mathrm{CN})_{4}\right]^{2-}$ is diamagnetic.

$$
\begin{align*}
& \text { (i) (a) . } 36 \\
& \text { ↔ A }  \tag{ii}\\
& \downarrow \tag{b}
\end{align*}
$$

$$
\begin{aligned}
& \text { - } 1.8 \times 10^{-5} \text { K }
\end{aligned}
$$

(a) (i) Write two differences between rate and rate constant of a reaction.
(ii) Derive integrated rate law for a zero order reaction $\mathrm{A} \rightarrow$ product.

## OR

(b) Find the pH of a buffer solution containing 0.20 mole per litre sodium acetate and 0.18 mole per litre acetic acid. $\mathrm{K}_{\mathrm{a}}$ for acetic acid is $1.8 \times 10^{-5}$.

(a) .37
(Benzoin) (i)


$\downarrow$

(i) (b)

(a) How will you convert benzaldehyde into the following compounds ?
(i) Benzoin
(ii) Cinnamic acid
(iii) Malachite green

## OR

(b) (i) Differentiate primary, secondary and tertiary alcohols using Lucas test.
(ii) Give the uses of diethyl ether.

$$
\begin{aligned}
& \text { (a) } .38 \\
& \downarrow
\end{aligned}
$$

$$
\begin{align*}
& \text { ج } \tag{b}
\end{align*}
$$

(a) Describe adsorption theory of catalysis.

## OR

(b) A compound ' A ' of molecular formula $\mathrm{C}_{2} \mathrm{H}_{3} \mathrm{~N}$ on reduction with $\mathrm{Na}(\mathrm{Hg}) / \mathrm{C}_{2} \mathrm{H}_{5} \mathrm{OH}$ gives ' B ' of molecular formula $\mathrm{C}_{2} \mathrm{H}_{7} \mathrm{~N}$ which undergoes carbylamine test. Compound ' B ' on reaction with nitrous acid gives compound ' C ' of molecular formula $\mathrm{C}_{2} \mathrm{H}_{6} \mathrm{O}$ by liberating nitrogen. Identify $\mathrm{A}, \mathrm{B}$ and C and write the reactions involved.
$\left.\begin{array}{r}\text { பதிவு எண் } \\ \text { Register Number }\end{array}\right)$

## PART - III

## உயிாியல் / BIOLOGY

( தமிழ் மற்றும் ஆங்கில வழி / Tamil \& English Version )

கால அளவு : 3.00 மணி நேரம் ]
Time Allowed : 3.00 Hours ]
[ மொத்த மதிப்பெண்கள் : 70
[Maximum Marks : 70

அறிவுரைகள் : (1) அனைத்து வினாக்களும் சாியாக பதிவாகி உள்ளதா என்பதனை சாிபார்த்துக் கொள்ளவும். அச்சுப்பதிவில் குறையிருப்பின் அறைக் கண்காணிப்பாளரிடம் உடனடியாகத் தொிவிக்கவும்.
(2) நீலம் அல்லது கருப்பு மையினை மட்டுமே எழுதுவதற்குப் பயன்படுத்த வேண்டும். படங்கள் வரைவதற்கும், அடிக்கோடிடுவதற்கும் பென்சில் பயன்படித்தவும்.

Instructions : (1) Check the question paper for fairness of printing. If there is any lack of fairness, inform the Hall Supervisor immediately.
(2) Use Blue or Black ink to write and underline and pencil to draw diagrams.

குறிப்பு : பகுதி-I (உயிரி-தாவரவியல்), பகுதி-II (உயிரி-விலங்கியல்) தனித்தனி விடைத்தாளில் விடையளிக்கவும்.

Note : Candidate should answer Part-I (Bio-Botany) \& Part-II (Bio-Zoology) in separate answer-books.

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பகுதி - I (உயிாி-தாவரவியல்) / PART - I (BIO-BOTANY)
    (மதிப்பெண்கள் : 35) / (Marks : 35)
```

    பிாிவு - 1/SECTION - 1
    குறிப்பு : (i) அனைத்து வினாக்களுக்கும் விடையளிக்கவும்.
(ii) கொடுக்கப்பட்டுள்ள மாற்று விடைகளில் மிகவும் ஏற்புடைய விடையைத் தேர்ந்தெடுத்துக் குறியீட்டுடன் விடையினையும் சோ்த்து எழுதவும்.
Note: (i) Answer all the questions.
(ii) Choose the most appropriate answer from the given four alternatives and write the option code and the corresponding answer.

