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JEE
(Main)
PAPER-1 (B.E./B. TECH.)
2022


COMPUTER BASED TEST (CBT)
Memory Based Questions & Solutions

Date: 28 July, 2022 (SHIFT-2) | TIME : (3.00 p.m. to 6.00 p.m)
Duration: 3 Hours | Max. Marks: 300

SUBJECT: CHEMISTRY

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PART : CHEMISTRY

1. Identify total number of state variables from the following.
Heat, Volume, Enthalpy, Internal energy

Ans. (3)

Sol. State variable \Rightarrow Volume, Enthalpy, Internal energy,
Path function \Rightarrow Heat

2. Correct order of increasing metallic character of following element is :
Na, Mg, Be, Si, P
(1) $P < Si < Be < Mg < Na$ (2) $P < Si < Na < Be < Mg$
(3) $Na < Mg < Be < Si < P$ (4) $Be < Mg < Na < Si < P$
- Ans. (1)
Sol. On moving top to bottom metallic character increases and on moving left to right metallic character decreases.
3. 0.2 M, 2 L H_2SO_4 is mixed with 0.1 M 2 L NaCl, then molarity of Na_2SO_4 is _____ $\times 10^{-3}$.
- Ans. (25)
Sol. $H_2SO_4 + 2NaCl \longrightarrow Na_2SO_4$
Mole 0.4 0.2 LR is NaCl
0.3 0 0.1
Molarity of $Na_2SO_4 = \frac{0.1}{4} = 0.025$
4. Assertion : Metal ion from their compound easily reduce in molten state than in solid state.
Reason : ΔG become more negative due to increase in entropy.
(1) Assertion is true, Reason is true and Reason is correct explanation of Assertion
(2) Assertion is true, Reason is true and Reason is not correct explanation of Assertion.
(3) Assertion is true and Reason is false
(4) Assertion is false and Reason is true
- Ans. (1)
Sol. $\Delta G = \Delta H - T\Delta S$
So on melting entropy is increases and ΔG become more negative so metal ion get easily reduced.
5. When liquid A and liquid B are mixed from an ideal solution in which mole fraction of A in vapour phase is 0.5 while in liquid phase is 0.2. If total vapour pressure is 0.8 atm, then vapour pressure of pure A is _____ atm.
- Ans. (2)
Sol. $P_A = P_A^0 \times X_A = (P_{Total})Y_A$
(P_A^0) $0.2 = 0.8 \times 0.5$
 $P_A^0 = 2 \text{ atm}$

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6. Statement-I : An alloy of Li & Mg is used in aircraft construction.
Statement-II : Magnesium plays important roles in neuromuscular function, interneuronal transmission, cell membrane integrity and blood coagulation.
(1) Both Statement I and Statement II are correct
(2) Statement I is correct but Statement II is incorrect
(3) Statement I is incorrect but Statement II is correct.
(4) Both Statement I and Statement II are incorrect
- Ans. (4)
Sol. Statement-I : An alloy of Al & Mg is used in aircraft construction.
Statement-II : Calcium plays important roles in neuromuscular function, interneuronal transmission, cell membrane integrity and blood coagulation.
7. 2 mole of NO and 1 mole of O_2 on reaction gives 0.8 mole of NO_2 at equilibrium according to the reaction
 $2NO(g) + O_2(g) \longrightarrow 2NO_2(g)$
At 600 K, if equilibrium pressure is 1 atm, then find K_c
(Given $R = 0.0821 \frac{\text{atm} \times \text{lit}}{\text{mol} \times \text{K}}$) (Report your answer to nearest integer)
- Ans. (93)
Sol. $2NO(g) + O_2(g) \longrightarrow 2NO_2(g)$
t = 0 2 mole 1 mole 0
1.2 mole 0.6 mole 0.8

$$n_{\text{Total}} = 2.6$$

$$K_P = \frac{(P_{\text{NO}_2})^2}{(P_{\text{NO}})^2 (P_{\text{O}_2})} = \frac{\left(\frac{0.8}{2.6} \times 1\right)^2}{\left(\frac{1.2}{2.6}\right)^2 \left(\frac{0.6}{2.6}\right)} = \frac{(0.8)^2 \times 2.6}{(1.2)^2 \times 0.6} = \frac{1.664}{0.864} = 1.8824$$

$$K_P = K_C = (RT)^{\Delta n} \Rightarrow 1.8824 = K_C (0.082 \times 600)^{-1}$$

$$K_C = 92.6$$

8. Assertion : HCl is not used in permanganate titration.

Reason : Permanganate ion oxidise HCl to Cl_2 .

(1) Assertion is true, Reason is true and Reason is correct explanation of Assertion

(2) Assertion is true, Reason is true and Reason is not correct explanation of Assertion.

(3) Assertion is true and Reason is false

(4) Assertion is false and Reason is true

Ans. (1)

Sol. $\text{MnO}_4^- + \text{Cl}^- \longrightarrow \text{Mn}^{2+} + \text{Cl}_2(\text{g})$

KMnO_4 oxidise HCl to Cl_2 that's why for acidic medium HCl is not used in permanganate titration.

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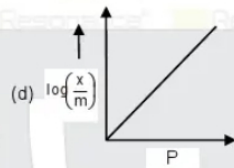
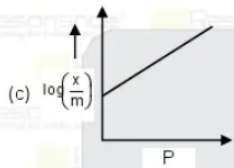
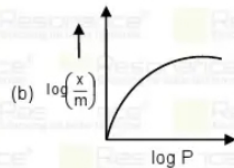
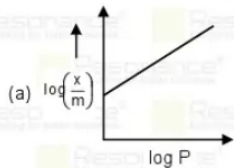
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9. How many of the following are not according to Freundlich adsorption isotherm.



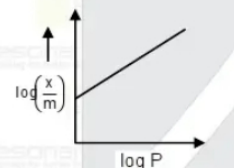
Ans. (3)

Sol. Freundlich adsorption isotherm.

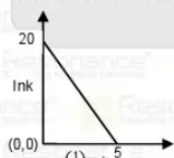
$$\frac{x}{m} = k(P)^{\frac{1}{n}}$$

$$\log\left(\frac{x}{m}\right) = \log k + \frac{1}{n} \log(P)$$

graph between $\log\left(\frac{x}{m}\right)$ Vs $\log(P)$



10. For a 1st order reaction following graph is obtained.



Using this graph find activation energy in calories. (Given $R = 2 \text{ Cal}$)
(Report your answer to nearest integer)

Ans. (8)

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Sol. $K = Ae^{-E_a/RT}$

$$\ln k = \ln A - \left(\frac{E_a}{R}\right) \frac{1}{T}$$

$$\text{Slope of graph} = -\left(\frac{E_a}{R}\right) = \left(\frac{0-20}{5-0}\right)$$

$$E_a = 4 \times 2 = 8 \text{ cal.}$$

11.

	List-I	List-II
	Complex	Hybridisation
(i)	$[\text{Ni}(\text{CO})_4]$	(a) dsp^2
(ii)	$[\text{Ni}(\text{CN})_4]^{2-}$	(b) sp^3
(iii)	$[\text{CoF}_6]^{3-}$	(c) d^2sp^3
(iv)	$[\text{Co}(\text{CN})_6]^{3-}$	(d) sp^3d^2

	(i)	(ii)	(iii)	(iv)
(1)	b	a	d	c
(2)	a	b	c	d
(3)	a	b	d	c
(4)	b	a	c	d

Ans. (1)

Sol.

	List-I	List-II
	Complex	Hybridisation
(i)	$[\text{Ni}(\text{CO})_4]$	(a) sp^3
(ii)	$[\text{Ni}(\text{CN})_4]^{2-}$	(b) dsp^2
(iii)	$[\text{CoF}_6]^{3-}$	(c) sp^3d^2
(iv)	$[\text{Co}(\text{CN})_6]^{3-}$	(d) d^2sp^3

12. 14 gram of Nitrogen present in closed container at 300 K. Find how much energy is given to system so that rms velocity of nitrogen molecule become double.

(Assume nitrogen act as ideal gas) $[R = 8.3 \text{ J/mol} \times K]$

(1) 9337.55 J (2) 3112.5 J (3) 12450 J (4) 6225 J

Ans. (1)

Sol.

$$U_{\text{rms}} = \sqrt{\frac{3RT}{M}}$$

$$\frac{(U_{\text{rms}})_2}{(U_{\text{rms}})_1} = \sqrt{\frac{T_2}{T_1}} = 2$$

$$\Rightarrow T_2 = 4T_1$$

$$T_2 = 1200 \text{ K}$$

$$\Delta U = \left(\frac{f}{2}\right) nR\Delta T$$

$$= \frac{5}{2} \times \frac{1}{2} \times 8.3 \times 900$$

$$= 9337.5 \text{ J}$$

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13. An element crystallize in FCC unit cell with edge length 4×10^{-8} cm and have density 9.03 gram/cm³, then molar mass of element is : (Given : $N_A = 6.02 \times 10^{23}$)
(Report your answer to nearest integer)

Ans. (87)

Sol. $d = \frac{Z \times M}{N_A \times \text{Volume}}$

$$9.03 = \frac{4 \times M}{6.02 \times 10^{23} \times (4 \times 10^{-8})^3}$$

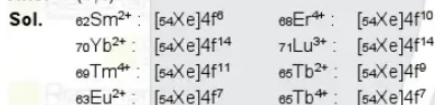
$$M = \frac{9.03 \times 6.02 \times 10^{23} \times 64 \times 10^{-24}}{4}$$

$$= 86.97 \text{ gram}$$

14. Which of the following are not isoelectronic species :

(1) Sm^{2+} , Er^{4+} (2) Yb^{2+} , Lu^{3+} (3) Tm^{4+} , Tb^{2+} (4) Eu^{2+} , Tb^{4+}

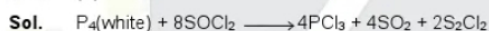
Ans. (1,4)



15. What products are formed in the reaction of white phosphorous with thionylchloride :

(1) PCl_3 , SO_2 , S_2Cl_2 (2) PCl_5 , SO_2 , S_2Cl_2
 (3) PCl_3 , SO_2 (4) PCl_3 , SO_2 , Cl_2

Ans. (1)



16. When concentrate HNO_3 react with I_2 product obtained are :

(1) HIO_3 , NO_2 , H_2O (2) HIO_4 , NO_2 , H_2O
 (3) HIO_3 , NO_2 , H_2O (4) HIO_4 , N_2O , H_2O

Ans. (1)



17. Assertion : Zero overlap is a type of out of phase overlap.

Reason : It is due to different orientation direction of approach

(1) Assertion is true, Reason is true and Reason is correct explanation of Assertion
 (2) Assertion is true, Reason is true and Reason is not correct explanation of Assertion.
 (3) Assertion is true and Reason is false
 (4) Assertion is false and Reason is true

Ans. (1)

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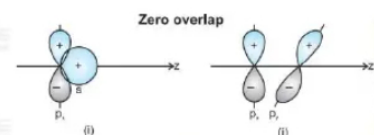
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Sol.



Zero overlap (out of phase due to different orientation direction of approach)

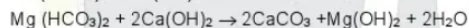
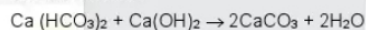
40. Is removal of temporary hardness using Clark's method. Product obtained are :

10. In removal of temporary hard using Clark's method. Product obtained are :

- (1) $\text{CaCO}_3, \text{Mg(OH)}_2$ (2) $\text{CaCO}_3, \text{MgCO}_3$
 (3) $\text{Ca(OH)}_2, \text{MgCO}_3$ (4) $\text{Ca(OH)}_2, \text{Mg(OH)}_2$

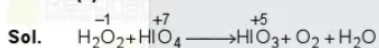
Ans. (1)

Sol. Clark's method :



19. H_2O_2 can be oxidise by powerful oxidising agent HIO_4 and O_2 is obtained.
 Find oxidation state of iodine in the product.

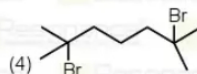
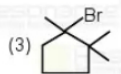
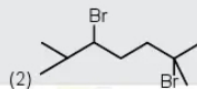
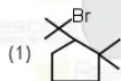
Ans. (5)



20. Nitrogen is not react with oxygen in atmosphere due to :

- (1) Nitrogen is unreactive in atmospheric condition
 (2) Nitrogen oxide are unstable
 (3) The reaction highly endothermic
 (4) The reaction is highly exothermic

Ans. (3)







Ans. (1)

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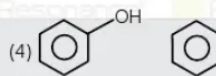
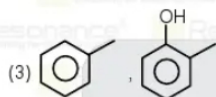
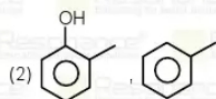
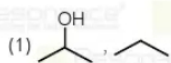
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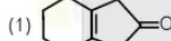
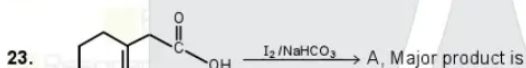
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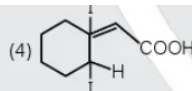
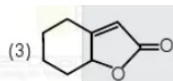


A and B are :



Ans. (4)





Ans. (3)

24. Statement-I : Aniline on nitration give ortho, meta, para derivative of aniline.
Statement-II : Nitration of aniline is carried out in highly acidic solution.
(1) Statement-I is correct only.
(2) Statement-II is correct only.
(3) Statement-I is correct and Statement-II is incorrect.
(4) Statement-I is incorrect and Statement-II is correct.

Ans. (3)

25. Statement I : Thin layer chromatography involves separation of substances based on adsorption.
Statement II : Silica gel coated on glass plate in thin layer chromatography is used as an adsorbent.
(1) Statement-I is correct only.
(2) Statement-II is correct only.
(3) Statement-I and Statement-II both are correct.
(4) Statement-I is incorrect and Statement-II is correct.

Ans. (3)

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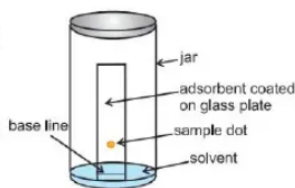
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Sol.



$$R_f = \frac{\text{Distance moved by the substance from base line (x)}}{\text{Distance moved by the solvent from base line (y)}}$$

26. Fructose $\xrightarrow{\text{HCN/H}_3\text{O}^+} \text{P}$
Fructose $\xrightarrow{\text{Red P/HI}} \text{Q}$
P and Q are respectively.
(1) $\text{C}_7\text{H}_{14}\text{O}_8$, C_6H_{14} (2) $\text{C}_7\text{H}_{13}\text{O}_8$, $\text{C}_6\text{H}_{14}\text{O}$ (3) $\text{C}_7\text{H}_{12}\text{O}_8$, C_6H_{14} (4) $\text{C}_7\text{H}_{14}\text{O}_8$, $\text{C}_6\text{H}_{14}\text{O}_8$

Ans. (1)

- Sol. Fructose $\xrightarrow{\text{HCN/H}_3\text{O}^+} \text{C}_7\text{H}_{14}\text{O}_8$
Fructose $\xrightarrow{\text{Red P/HI}} \text{C}_6\text{H}_{14}$

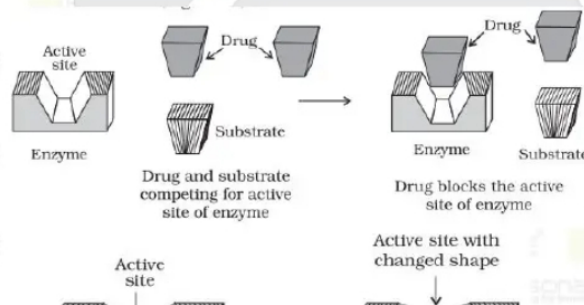
27. Statement I : The only way by which an inhibitor can prevent action of enzyme is by blocking the active site.

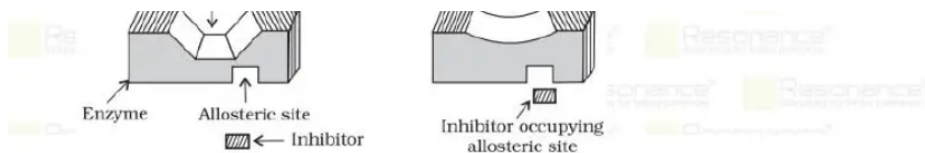
Statement II : Inhibitor can form a strong covalent bond with enzyme.

- (1) Statement-I is correct only.
(2) Statement-II is correct only.
(3) Statement-I is correct and Statement-II is incorrect.
(4) Statement-I is incorrect and Statement-II is correct.

Ans. (4)

Sol.





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28. Match the following :

(A) Elastomers	(P) $\left[\text{CH}_2 - \underset{\text{Cl}}{\text{CH}} \right]_n$
(B) Fibres	(Q) $\left[\text{CH}_2 - \underset{\text{Cl}}{\text{C}} = \text{CH} - \text{CH}_2 \right]_n$
(C) Thermoplastic polymers	(R) $\left[\text{H} - \text{N} - (\text{CH}_2)_6 - \text{N} - \text{C}(=\text{O}) - (\text{CH}_2)_4 - \text{C}(=\text{O}) \right]_n$
(D) Thermosetting polymers	(S) $\left[\text{C}_6\text{H}_3(\text{OH})_2 - \text{CH}_2 - \text{C}_6\text{H}_3(\text{OH})_2 - \text{CH}_2 \right]_n$

(1) (A)-Q, (B)-R, (C)-P, (D)-S

(2) (A)-S, (B)-R, (C)-Q, (D)-P

(3) (A)-R, (B)-Q, (C)-P, (D)-S

(4) (A)-P, (B)-Q, (C)-S, (D)-R

Ans.

Sol.

(A) Elastomers	(P) $\left[\text{CH}_2 - \underset{\text{Cl}}{\text{C}} = \text{CH} - \text{CH}_2 \right]_n$
(B) Fibres	(Q) $\left[\text{H} - \text{N} - (\text{CH}_2)_6 - \text{N} - \text{C}(=\text{O}) - (\text{CH}_2)_4 - \text{C}(=\text{O}) \right]_n$
(C) Thermoplastic polymers	(R) $\left[\text{CH}_2 - \underset{\text{Cl}}{\text{CH}} \right]_n$
(D) Thermosetting polymers	(S) $\left[\text{C}_6\text{H}_3(\text{OH})_2 - \text{CH}_2 - \text{C}_6\text{H}_3(\text{OH})_2 - \text{CH}_2 \right]_n$

29. The correct order of rate of nitration of benzene with following compound is :

(i) P-Xylene (ii) Bromobenzene (iii) Mesitylene (iv) Nitrobenzene (v) Benzene

(1) (iii) > (i) > (iv) > (ii) > (v)

(2) (i) > (iii) > (iv) > (ii) > (v)

(3) (iii) > (iv) > (i) > (ii) > (v)

(4) (iii) > (i) > (iv) > (v) > (ii)

Ans.

(1)

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